

Experiment 14

Data and Calculations: Heat Effects and Calorimetry

A. Specific Heat

	Trial 1		Trial 2
Mass of stoppered test tube plus metal	<u>45.393</u> g	→	<u>45.406</u> g
Mass of test tube and stopper	<u>22.525</u> g	→	<u>22.539</u> g
Mass of calorimeter	<u>3.978</u> <u>47.173</u> g	→	<u>8.608</u> g
Mass of calorimeter and water	<u>45.798</u> g		<u>48.295</u> g
Mass of water	<u>39.175</u> g		<u>39.687</u> g
Mass of metal	<u>22.868</u> g	→	<u>22.868</u> g
Initial temperature of water in calorimeter	<u>21.2</u> °C		<u>21.9</u> °C
Initial temperature of metal (assume 100°C unless directed to do otherwise)	<u>99.8</u> °C	→	<u>100.3</u> °C
Equilibrium temperature of metal and water in calorimeter	<u>25.3</u> °C		<u>24.5</u> °C
$\Delta t_{\text{water}} (t_{\text{final}} - t_{\text{initial}})$	_____ °C		_____ °C
Δt_{metal}	_____ °C		_____ °C
$q_{\text{H}_2\text{O}}$	_____ J		_____ J
Specific heat of the metal (Eq. 3)	_____ J/g°C		_____ J/g°C
Approximate molar mass of metal	_____		_____
Unknown metal #			<u>26Y</u>

B. Heat of Solution

Mass of calorimeter plus water	<u>56.898</u> g
Mass of beaker	<u>29.309</u> g

(continued on following page)

Lab Manual Chemistry Heat Effects And Calorimetry

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Table of Contents Lab Manual Chemistry Heat Effects And Calorimetry

1. Understanding the eBook Lab Manual Chemistry Heat Effects And Calorimetry
 - The Rise of Digital Reading Lab Manual Chemistry Heat Effects And Calorimetry
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab Manual Chemistry Heat Effects And Calorimetry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab Manual Chemistry Heat Effects And Calorimetry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab Manual Chemistry Heat Effects And Calorimetry
 - Personalized Recommendations
 - Lab Manual Chemistry Heat Effects And Calorimetry User Reviews and Ratings
 - Lab Manual Chemistry Heat Effects And Calorimetry and Bestseller Lists
5. Accessing Lab Manual Chemistry Heat Effects And Calorimetry Free and Paid eBooks
 - Lab Manual Chemistry Heat Effects And Calorimetry Public Domain eBooks
 - Lab Manual Chemistry Heat Effects And Calorimetry eBook Subscription Services
 - Lab Manual Chemistry Heat Effects And Calorimetry Budget-Friendly Options
6. Navigating Lab Manual Chemistry Heat Effects And Calorimetry eBook Formats

- ePub, PDF, MOBI, and More
 - Lab Manual Chemistry Heat Effects And Calorimetry Compatibility with Devices
 - Lab Manual Chemistry Heat Effects And Calorimetry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab Manual Chemistry Heat Effects And Calorimetry
 - Highlighting and Note-Taking Lab Manual Chemistry Heat Effects And Calorimetry
 - Interactive Elements Lab Manual Chemistry Heat Effects And Calorimetry
 8. Staying Engaged with Lab Manual Chemistry Heat Effects And Calorimetry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab Manual Chemistry Heat Effects And Calorimetry
 9. Balancing eBooks and Physical Books Lab Manual Chemistry Heat Effects And Calorimetry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab Manual Chemistry Heat Effects And Calorimetry
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab Manual Chemistry Heat Effects And Calorimetry
 - Setting Reading Goals Lab Manual Chemistry Heat Effects And Calorimetry
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab Manual Chemistry Heat Effects And Calorimetry
 - Fact-Checking eBook Content of Lab Manual Chemistry Heat Effects And Calorimetry
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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