

Chemistry 100: Final Exam Questions

Question 1: Which compound is a substituted alkane?
 a. ethane, CH_3CH_3
 b. propene, $\text{CH}_3\text{CH}=\text{CH}_2$
 c. butanol, $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$
 d. acetone, CH_3COCH_3
 e. benzene, C_6H_6

Question 2: Which of the following is the correct Lewis structure for water?
 a. $\text{H}-\text{O}-\text{H}$
 b. $\text{H}-\text{O}-\text{H}$ with two lone pairs on oxygen
 c. $\text{H}-\text{O}-\text{H}$ with one lone pair on oxygen
 d. $\text{H}-\text{O}-\text{H}$ with three lone pairs on oxygen

Question 3: Which of the following represents the correct Lewis structure for carbon dioxide?
 a. $\text{O}=\text{C}=\text{O}$
 b. $\text{O}=\text{C}=\text{O}$ with four lone pairs on each oxygen
 c. $\text{O}=\text{C}=\text{O}$ with two lone pairs on each oxygen
 d. $\text{O}=\text{C}=\text{O}$ with one lone pair on each oxygen

Question 4: A gas sample weighs 0.450 g. It has a volume of 0.100 L at 0.987 atm and 25.0°C. What is the molar mass of the gas?
 a. 44.0 g/mol
 b. 44.0 g/mol
 c. 44.0 g/mol
 d. 44.0 g/mol
 e. 44.0 g/mol

Question 5: The compound with the empirical formula CH_2O has a molar mass of 60.0 g/mol. What is the molecular formula?
 a. CH_2O
 b. $\text{C}_2\text{H}_4\text{O}_2$
 c. $\text{C}_3\text{H}_6\text{O}_3$
 d. $\text{C}_4\text{H}_8\text{O}_4$
 e. $\text{C}_5\text{H}_{10}\text{O}_5$

Question 6: A sample of pure carbon dioxide gas is collected over water at 25.0°C. The total pressure is 760 mmHg. What is the partial pressure of the carbon dioxide?
 a. 760 mmHg
 b. 750 mmHg
 c. 740 mmHg
 d. 730 mmHg
 e. 720 mmHg

Question 7: Which of the following represents the correct Lewis structure for ammonia?
 a. $\text{H}-\text{N}-\text{H}$
 b. $\text{H}-\text{N}-\text{H}$ with one lone pair on nitrogen
 c. $\text{H}-\text{N}-\text{H}$ with two lone pairs on nitrogen
 d. $\text{H}-\text{N}-\text{H}$ with three lone pairs on nitrogen
 e. $\text{H}-\text{N}-\text{H}$ with four lone pairs on nitrogen

Question 8: The density of a gas is 1.25 g/L at 0.987 atm and 25.0°C. What is the molar mass of the gas?
 a. 44.0 g/mol
 b. 44.0 g/mol
 c. 44.0 g/mol
 d. 44.0 g/mol
 e. 44.0 g/mol

Question 9: The density of a gas is 1.25 g/L at 0.987 atm and 25.0°C. What is the molar mass of the gas?
 a. 44.0 g/mol
 b. 44.0 g/mol
 c. 44.0 g/mol
 d. 44.0 g/mol
 e. 44.0 g/mol

Question 10: A gas sample weighs 0.450 g. It has a volume of 0.100 L at 0.987 atm and 25.0°C. What is the molar mass of the gas?
 a. 44.0 g/mol
 b. 44.0 g/mol
 c. 44.0 g/mol
 d. 44.0 g/mol
 e. 44.0 g/mol

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