

**Finishing Part A**  
Use the chart below to determine the pH of each solution. Each pH indicator changes color based on the pH of the solution. Look at Methyl Red as an example. Methyl red will be red at and below a pH of 4.8. Additionally, methyl red will be yellow at or above a pH of 6.0. In between a pH of 4.8 and 6.0, methyl red will be orange (red and yellow make orange).

Each pH indicator has a limited range. Use the combination of the indicators to narrow down the pH range for each solution. For example, consider a solution with unknown pH. Start with the indicators that change color around a pH of 7.0, which will be bromothymol blue and phenol red. Our bromothymol blue indicator is blue, so the pH must be greater than 7.6. Likewise, the phenol red indicator is red, so the pH must be greater than 8.0. However, our phenolphthalein solution is colorless, indicating a pH lower than 8.2. Therefore, we have determined that the solution must have a pH between 8.0 and 8.2.

**Abstract**

1. Record the experimental pH range for each solution below

Distilled Water	
Tap Water	
0.1M NaOH	
0.1M $\text{NH}_4\text{OH}$	
0.1M HAc	
0.1M HCl	
0.1M $\text{H}_2\text{SO}_4$	

2. Calculate the theoretical pH for each solution and record it below.
- Distilled Water

# Gsu Chemistry 1212 Lab Manual

**Mark A. Griep**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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