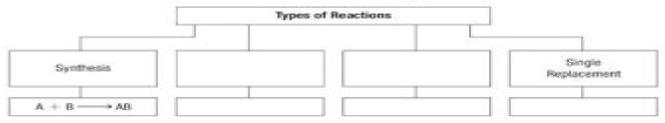
## Section 7.2 Types of Reactions

(pages 199-205)

This section discusses how chemical reactions are classified into different types.

#### Reading Strategy (page 199)

Previewing Skim the section and begin a concept map like the one below that identifies types of reactions with a general form. As you read, add the general form of each type of reaction. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook.



#### Classifying Reactions (pages 199-204)

- 2. Circle the letter of each equation that represents a synthesis reaction.
  - a.  $2Na + Cl_2 \longrightarrow 2NaCl$
  - b. 2NaCl --- 2Na + Cla
  - c.  $2H_2O \longrightarrow 2H_2 + O_2$
  - $d. 2H_2 + O_2 \longrightarrow 2H_2O$
- Is the following sentence true or false? A decomposition reaction is the opposite of a synthesis reaction.
- Write the equation for the decomposition of calcium carbonate into calcium oxide and carbon dioxide.
- 5. Circle the letter of the correct answer. Copper reacts with silver nitrate in a single-replacement reaction. What are the products of this reaction?
  - a. copper(II) nitride and silver oxide
  - b. copper(II) nitrate and silver
  - c. copper(II) oxide and silver nitrate
  - d. copper, nitrogen, and silver oxide

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## Table of Contents Guided Reading And Study Workbook Chapter 7

- 1. Understanding the eBook Guided Reading And Study Workbook Chapter 7
  - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 7
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Guided Reading And Study Workbook Chapter 7
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - $\circ\,$  Features to Look for in an Guided Reading And Study Workbook Chapter 7
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 7
  - Personalized Recommendations
  - Guided Reading And Study Workbook Chapter 7 User Reviews and Ratings
  - Guided Reading And Study Workbook Chapter 7 and Bestseller Lists
- 5. Accessing Guided Reading And Study Workbook Chapter 7 Free and Paid eBooks

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  - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 7
  - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 7
  - o Interactive Elements Guided Reading And Study Workbook Chapter 7
- 8. Staying Engaged with Guided Reading And Study Workbook Chapter 7
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Reading And Study Workbook Chapter 7
- 9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 7
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 7
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 7
  - Setting Reading Goals Guided Reading And Study Workbook Chapter 7
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 7
  - Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 7
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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