

Name _____ Date _____ Class _____

CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.3 Limiting Reactants

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 25–27 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and State of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	8 molecules	1, 4 molecules	2, O_2	3, 4 molecules NO
1.00 mol	2.00 mol	4, 2.00 mol	5, none	6, none
4.00 mol	4.00 mol	7, 4.00 mol	8, NO	9, 2.00 mol O_2
1.00 mol	1.00 mol	10, 7.00 mol	11, NO	12, 1.50 mol O_2
1.00 mol	4.00 mol	13, 2.00 mol	14, O_2	15, 2.00 mol NO
0.500 mol	0.200 mol	16, 0.200 mol	17, NO	18, 0.400 mol O_2
32.00 g	60.02 g	19, 62.02 g	20, none	21, none
16.00 g	60.02 g	22, 46.01 g	23, O_2	24, 58.12 g NO
16.00 g	20.00 g	25, 28.76 g	26, O_2	27, 1.24 g NO

balanced equation mole ratio = 2 mol NO/1 mol O_2

$16.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.5125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.5125 \text{ mol } \text{O}_2 = 2.132 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of NO: O_2 is larger than the balanced equation mole ratio of NO: O_2 , there is an excess of NO; O_2 is the limiting reactant.

Mass of NO used = $0.5125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.4250 \text{ mol } \text{NO}$

$0.4250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 12.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.4250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 19.56 \text{ g } \text{NO}_2$

Excess NO = $20.00 \text{ g } \text{NO} - 12.76 \text{ g } \text{NO} = 7.24 \text{ g } \text{NO}$

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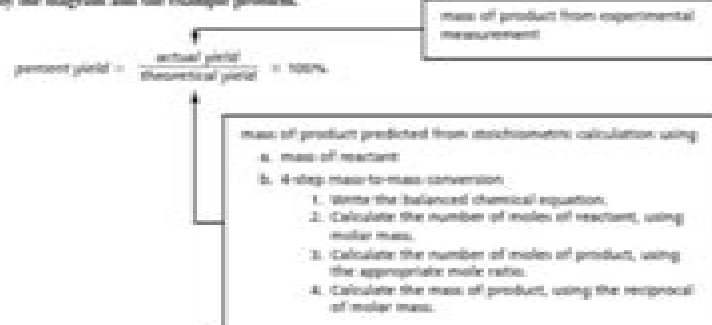
CHAPTER 12

STUDY GUIDE FOR CONTENT MASTERY

Section 12.4 Percent Yield

In your textbook, read about the yield of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of gallium oxide, a substance used in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the oxide from a 7.00-g sample of gallium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 187.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield: unknown $7.42 \text{ g } \text{Ga}_2\text{O}_3$
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 7.00 \text{ g } \text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}} = 0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
- theoretical yield: $0.100 \text{ mol } \text{Ga}_2\text{O}_3 \times \frac{187.44 \text{ g } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}_2\text{O}_3} = 18.74 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } \text{Ga}_2\text{O}_3}{18.74 \text{ g } \text{Ga}_2\text{O}_3} \times 100 = 7.42 \text{ g } \text{Ga}_2\text{O}_3 / 18.74 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

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G Psacharopoulos



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Table of Contents Limiting Reactants Study Guide

1. Understanding the eBook Limiting Reactants Study Guide
 - The Rise of Digital Reading Limiting Reactants Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reactants Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reactants Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reactants Study Guide
 - Personalized Recommendations
 - Limiting Reactants Study Guide User Reviews and Ratings
 - Limiting Reactants Study Guide and Bestseller Lists
5. Accessing Limiting Reactants Study Guide Free and Paid eBooks
 - Limiting Reactants Study Guide Public Domain eBooks
 - Limiting Reactants Study Guide eBook Subscription Services
 - Limiting Reactants Study Guide Budget-Friendly Options

6. Navigating Limiting Reactants Study Guide eBook Formats
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 - Limiting Reactants Study Guide Compatibility with Devices
 - Limiting Reactants Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reactants Study Guide
 - Highlighting and Note-Taking Limiting Reactants Study Guide
 - Interactive Elements Limiting Reactants Study Guide
8. Staying Engaged with Limiting Reactants Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reactants Study Guide
9. Balancing eBooks and Physical Books Limiting Reactants Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reactants Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting Reactants Study Guide
 - Setting Reading Goals Limiting Reactants Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting Reactants Study Guide
 - Fact-Checking eBook Content of Limiting Reactants Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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