

MODUL BELAJAR KIMIA 8

Kompetensi Dasar : 1.1	Memahami Struktur atom,Sifat-sifat Unsur dan Ikatan kimia	Periodik
Materi Pokok :	1. Ikatan kovalen Polar dan Non Polar 2. Ikatan Logam	
Waktu :	2 x 45 menit	

1. IKATAN KOVALEN POLAR

Atom-atom pembentuknya mempunyai gaya tarik yang tidak sama terhadap pasangan elektron persekutuan. Hal ini terjadi karena beda keelektronegatifan kedua atomnya. Elektron persekutuan akan bergeser ke arah atom yang lebih elektronegatif akibatnya terjadi pemisahan kutub positif dan negatif. Dalam senyawa HCl ini, Cl mempunyai keelektronegatifan yang lebih besar dari H, sehingga pasangan elektron lebih tertarik ke arah Cl, akibatnya H relatif lebih elektropositif sedangkan Cl relatif menjadi elektronegatif.

Pemisahan muatan ini menjadikan molekul itu bersifat polar dan memiliki "momen dipol" sebesar:

$$T = q \cdot l$$

dimana :

T = momen dipol

q = kelebihan muatan pada masing-masing atom

l = jarak antara kedua inti atom

2. IKATAN KOVALEN NON POLAR

Tidak muatan negatif elektron persekutuan berhimpit, sehingga pada molekul pembentuknya tidak terjadi momen dipol, dengan perkataan lain bahwa elektron persekutuan mendapat gaya tarik yang sama. Contoh: Kedua atom H mempunyai harga keelektronegatifan yang sama. Karena arah tarikan simetris, maka tidak muatan negatif elektron persekutuan berhimpit. Contoh lain adalah senyawa CO_2 , O_2 , Br_2 dan lain-lain

LATIHAN

1. Tentukan molekul berikut ini tentukan yang ikatannya polar dan yang non polar

a. CO_2

c. F_2

b. Br_2

d. PCl_3

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CL Gary



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