

Acids and bases

Chapter outline

- Recall the definitions of Brønsted–Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid–base pairs.
- Recall the terms ‘amphoteric’ and ‘amphiprotic’.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction. **HL**
- Deduce the dissociation constant for weak acids and bases. **HL**
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the pK_a or pK_b value of a weak acid or base. **HL**
- Explain how an acid–base indicator works and select a suitable indicator for a titration. **HL**

KEY TERMS

Acid deposition: A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth’s surface.

Brønsted–Lowry acid: A proton (H^+) donor.

Brønsted–Lowry base: A proton (H^+) acceptor.

pH: A measure of the concentration of hydrogen ions in an aqueous solution.
 $pH = -\log_{10}[H^+]$

Strong acid/base: An acid/base that is completely dissociated into its ions in aqueous solution.

Weak acid/base: An acid/base that is only partially dissociated into its ions in aqueous solution.

Guided Study Workbook Acids And Bases Answers

Pamela Mills,Amina El-Ashmawy

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Organic Chemistry Study Guide with Solutions Manual Neil E. Schore, 2007 The guide includes chapter introductions that highlight new material chapter outlines detailed comments for each chapter section a glossary and solutions to the end of chapter problems presented in a way that shows students how to reason their way to the answer

Study Guide and Solutions Manual Neil E. Schore, 2002-08-02
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chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

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