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web solution 1 the two half reactions balanced as if in acidic solution $2\text{CN}^{-} + \text{Au} \rightarrow \text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$ $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ make electrons equal convert to basic solution $4\text{CN}^{-} + 2\text{Au} \rightarrow 2\text{Au}(\text{CN})_2^{-} + 2\text{e}^{-}$ multiplied by a factor of 2 $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ the final answer

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the loss of electrons resulting in a more positively charged species reduction the gain of electrons resulting in a

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