

Chapter 8 Chemical Reactions Chapter Review Answers

1. The evolution of energy, the production of a gas, the formation of a precipitate, and a color change.
2. A correctly written chemical equation must represent the known facts, contain the correct formulas for the reactants and products, and satisfy the law of conservation of mass.
3.
 - a. It is a number that appears in front of a formula in a chemical equation.
 - b. The number of atoms of each type in the formula is multiplied by the coefficient.
4. methane + oxygen $\rightarrow$ carbon dioxide + water;
 $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$;
 $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$
5. A chemical equation provides the number of molecules, moles, and atoms of each element or compound composing the reactants and products as well as the mole ratios of the reactants to each other and to the products.
6. Word and formula equations have only qualitative meaning; they do not express the relative quantities of a reaction's reactants or products.
7.
 - a. a solution in water
 - b. a substance that accelerates a chemical reaction but can be recovered unchanged
 - c. a chemical reaction in which the products re-form the original reactants
8.
 - a. KOH
 - b. $\text{Ca}(\text{NO}_3)_2$
 - c. Na_2CO_3
 - d. CCl_4
 - e. MgBr_2
9. (1) Balance different types of atoms one at a time; (2) balance types of atoms that appear only once on each side of the equation first; (3) balance as single units any polyatomic ions that appear on both sides of the equation; and (4) balance H atoms and O atoms last.
10.
 - a. 6N
 - b. 4H, 2O
 - c. 4H, 4N, 12O
 - d. 2Ca, 4O, 4H
 - e. 3Ba, 6Cl, 18O
 - f. 5Fe, 10N, 30O
 - g. 12Mg, 8P, 32O
 - h. 4N, 16H, 2S, 8O
 - i. 12Al, 188Se, 72O
 - j. 12C, 32H
11.
 - a. $2\text{ZnS}(\text{s}) + 3\text{O}_2(\text{g}) \rightarrow 2\text{ZnO}(\text{s}) + 2\text{SO}_2(\text{g})$
 - b. $2\text{HCl}(\text{aq}) + \text{Ba}(\text{OH})_2(\text{aq}) \rightarrow \text{BaCl}_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
 - c. $2\text{HNO}_3(\text{aq}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}(\text{NO}_3)_2(\text{aq}) + 2\text{H}_2\text{O}(\text{l})$
12.
 - a. Solid zinc sulfide reacts with oxygen gas to form solid zinc oxide and sulfur dioxide gas.
 - b. When solid calcium hydride is added to water, aqueous calcium hydroxide and hydrogen gas are formed.
 - c. Aqueous silver nitrate mixed with aqueous potassium iodide produces a precipitate of silver iodide and aqueous potassium nitrate.
13.
 - a. $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - b. $2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$
 - c. $\text{Pb}(\text{CH}_3\text{COO})_2 + \text{H}_2\text{S} \rightarrow \text{PbS} + 2\text{CH}_3\text{COOH}$
14.
 - a. LiO_2 is an incorrect formula; $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$
 - b. H_2Cl_2 is an incorrect formula; $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$
 - c. MgO_2 is an incorrect formula, and the equation as written is not balanced; $\text{MgCO}_3 \rightarrow \text{MgO} + \text{CO}_2$
 - d. I is an incorrect formula for iodine, and the equation is not balanced; $2\text{NaI} + \text{Cl}_2 \rightarrow 2\text{NaCl} + \text{I}_2$
15.
 - a. $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$
 - b. $\text{P}_2\text{O}_{10} + 6\text{H}_2\text{O} \rightarrow 4\text{H}_3\text{PO}_4$
 - c. $\text{Fe}_2\text{O}_3 + 3\text{CO} \rightarrow 2\text{Fe} + 3\text{CO}_2$
16. $\text{CH}_4(\text{g}) + 4\text{Cl}_2(\text{g}) \rightarrow \text{CCl}_4(\text{l}) + 4\text{HCl}(\text{g})$
17.
 - a. $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$
 - b. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$
 - c. $2\text{Li} + \text{Cl}_2 \rightarrow 2\text{LiCl}$
 - d. $\text{Ca} + \text{I}_2 \rightarrow \text{CaI}_2$
18.
 - (1) In a synthesis reaction, two or more substances combine to form a new substance: $\text{A} + \text{X} \rightarrow \text{AX}$
 - (2) In a decomposition reaction, a single compound undergoes a reaction that produces two or more simpler substances: $\text{AX} \rightarrow \text{A} + \text{X}$
 - (3) In a single-displacement reaction, one element replaces another element in a compound:
 $\text{A} + \text{BX} \rightarrow \text{AX} + \text{B}$ or $\text{Y} + \text{RX} \rightarrow \text{RY} + \text{X}$

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Physical Chemistry of Metallurgical Processes, Second Edition Mohammad Shamsuddin, 2021-06-22 This updated second edition retains its classroom tested treatment of physical chemistry of metallurgical topics such as roasting of sulfide minerals matte smelting converting structure properties and theories of slag reduction of oxides and reduction smelting interfacial phenomena steelmaking secondary steelmaking role of halides in extraction of metals refining hydrometallurgy and electrometallurgy and adds new data in worked out examples as well as up to date references to the literature. The book further explains the physical chemistry of various metallurgical topics steps involved in extraction of metals such as roasting matte smelting converting reduction smelting steelmaking reactions deoxidation stainless steelmaking vacuum degassing refining leaching chemical precipitation ion exchange solvent extraction cementation gaseous reduction and electrowinning. Each topic is illustrated with appropriate examples of applications of the technique in extraction of some common reactive rare or refractory metal together with worked out problems explaining the principle of the operation. The problems require imagination and critical analyses and also encourage readers for creative application of thermodynamic data in metal extraction. Updates and condenses text throughout the book by sequential arrangement of paragraphs in different chapters. Maximizes readers understanding of the physicochemical principles involved in extraction production of common and rare reactive metals by pyro as well as hydrometallurgical routes. Reinforces concepts presented with worked examples in each chapter explaining the process steps. Explains the physical chemistry of various metallurgical steps such as roasting matte smelting converting and reduction smelting steelmaking aqueous processing etc in extraction of metals. Collects and uniformly presents scattered information on physicochemical principles of metal production from various books and journals.

Bulletin of the Atomic Scientists, 1958-01 The Bulletin of the Atomic Scientists is the premier public resource on scientific and technological developments that impact global security. Founded by Manhattan Project Scientists the Bulletin's iconic Doomsday Clock stimulates solutions for a safer world.

Handbook of Systems Biology Marian Walhout, Marc Vidal, Job Dekker, 2012-12-31 This book provides an entry point into Systems Biology for researchers in genetics molecular biology cell biology microbiology and biomedical science to understand the key concepts to expanding their work. Chapters organized around broader themes of Organelles and Organisms Systems Properties of Biological Processes Cellular Networks and Systems Biology and Disease.

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