



Which of the following substances is a reaction intermediate?

- A) N_2O_2 B) HO C) H_2O D) N_2O E) N_2O_2 and N_2O

15. The value of ΔH° for the reaction below is -126 kJ . _____ kJ is released when 2.00 mol of NaOH is formed in the reaction?



- A) 252 B) 63 C) 3.9 D) 7.8 E) -126

16. Estimate the enthalpy change for the reaction below from the average bond energies given. There are two C-Cl and two C-H bonds in CH_2Cl_2 . Remember that energy is absorbed when bonds are broken and released when they are formed.



Average Bond Energies

C-H $+413 \text{ kJ/mol}$ Cl-Cl $+242 \text{ kJ/mol}$

H-Cl $+432 \text{ kJ/mol}$ C-Cl $+339 \text{ kJ/mol}$

- A) -232 kJ/mol B) $+578 \text{ kJ/mol}$ C) $+232 \text{ kJ/mol}$
D) -578 kJ/mol E) $+541 \text{ kJ/mol}$

17. According to the reaction profile below, what is ΔH for the reaction: $2\text{H}_2\text{O}(g) + 2\text{Br}_2(g) \rightarrow 4\text{HBr}(g) + \text{O}_2(g)$?



- A) -434 kJ B) 434 kJ C) -276 kJ D) 276 kJ E) 158 kJ

18. Given the following at ΔH° (25°C and 1.00 atm):



Calculate the ΔH° for the reaction below at 25°C:



- A) $+11.0 \text{ kJ}$ B) $+44.3 \text{ kJ}$ C) $+55.3 \text{ kJ}$ D) -22.1 kJ E) -55.3 kJ

19. Consider the following PE diagram for a catalyzed and uncatalyzed reaction:

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