

CK-12 Chemistry - Basic Answer Key

Chapter 4: Introduction to Chemistry

.1 Evolution of the Atomic Model

Check Your Understanding

1. What are the general properties of matter?

Answer: Matter can be described in terms of its physical characteristics as well as its chemical characteristics. For instance, the physical characteristics of liquid water are that it has a density of 1 g/mL. Many substances will dissolve readily in water – another physical characteristic.

2. We widely accept that all matter is comprised of similar kinds of particles that are too small to detect with the unaided eye. For example, we know that when a candle burns it gives off carbon dioxide and water vapor. Yet we cannot actually “see” these gases.

How can we study something we cannot directly see, but know exists? Can you give an example?

Answer: Even though we cannot readily see the carbon dioxide and water vapor produced, we could measure the mass changes in the candle. This would indirectly tell us how much gas was produced. We could also place a plate of glass over the burning candle to condense the water vapor which would then be visible in its liquid form. Chemistry depends on indirect measurement quite frequently to make sense and describe the chemical world.

Review Questions

Questions:

1. Compare and contrast the continuous theory of matter with the discontinuous theory of matter.

2. Which of the following did Proust discover?

- Elements combine in unpredictable ratios with other elements.
- The mass of an element is always changing even when combined with another element.
- Chemical compounds always contain the same proportion of elements by mass, regardless of amount.

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 - The Rise of Digital Reading Mastering Chemistry Answer Key Chapter 4 Download Free
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 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
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 - Personalized Recommendations

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 - Minimizing Distractions
 - Managing Screen Time
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