

Assessment 3

Chapter Test B**Chapter: Chemical Equations and Reactions**

Part 1 In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- _____ 1. The precipitation of a slightly soluble solid compound from a solution of two ionic reactants is called the formation of a
a. gas.
b. precipitate.
c. combustion reaction.
d. halogen.
- _____ 2. To balance a chemical equation, it is permissible to adjust the
a. coefficients.
b. subscripts.
c. formulae of the products.
d. number of products.
- _____ 3. In a chemical equation, the symbol (aq) indicates that the substance is
a. melted.
b. dissolved in water.
c. an acid.
d. insoluble.
- _____ 4. The tendency for a replacement reaction to occur depends on the
a. relative tendency of the two elements to lose electrons.
b. temperature difference.
c. relative electronegativity.
d. relative tendency of the two elements to gain electrons.
- _____ 5. The coefficients in a chemical equation
a. indicate the number of moles of each substance that react.
b. show the number of grams of each substance that react.
c. are the molar masses of the substances.
d. show the relative electronegativity of each atom.
- _____ 6. If metal X is more reactive than metal Y in the activity series, then
a. X will replace ions of Y in solution.
b. Y will replace ions of X in solution.
c. Y will displace oxygen only indirectly.
d. X will react with cold water.

Holt Chemistry Study Guide Answers Reaction Rate

Michal Rosen-Zvi



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Holt Chemistry Holt Rinehart & Winston, 2003-01-24 *Holt Chemistry* Ralph Thomas Myers, 2004 **ChemDiscovery Teacher Edition** Olga I. Agapova, 2002 **Books in Print Supplement**, 2002 **Children's Books in Print** R R Bowker Publishing, Bowker, 1999-12 Children's Books in Print, 2007, 2006 Economics and Thermodynamics Peter Burley, John Foster, 2013-03-09

Over the past two decades we have witnessed something of a revolution in the natural sciences as thermodynamic thinking evolved from an equilibrium or classical perspective to a nonequilibrium or self organisational one. In this transition thermodynamics has been applied in new ways and in new fields of inquiry. Chemical and biological evolutionary processes have been analysed increasingly in non equilibrium thermodynamical terms. Economics has since the late 19th century relied heavily upon metaphors and analogies derived from the natural sciences. Mechanical analogies cast in terms of traditional Newtonian physics and expressed in terms of Cartesian logic have been especially popular. Thermodynamics on the other hand has been less popular despite its early application in economics by Stanley Jevons the father of modern notions of utility maximisation in neoclassical economics and despite its promotion in economic contexts by Paul Samuelson the author of the definitive treatise upon which post war neoclassical economic theory was based namely his *Foundations of Economic Analysis*. The general neglect of thermodynamic thinking in economics was brought to our attention by Nicholas Georgescu Roegen in the late 1960s by which time economic theory evidenced in for example the Arrow Debreu general equilibrium system had become so sophisticated that it could not be penetrated by thermodynamical ideas. To Georgescu Roegen this presented something of a crisis in economics because neglect of thermodynamics led in his view to blindness amongst economists to an economy environment problem in the global economy.

Chapter Resource 2
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