



ANSWER GUIDE
PHYSICAL SCIENCES: GRADE 12
CLASS TEST: ACIDS & BASES

Written on: 25.06.2020

Total Marks: 65

1	1.1	B ✓✓	2
	1.2	C ✓✓	2
2	Hydrochloric acid is a strong acid and sodium hydroxide is a strong base, ✓ therefore the equivalence point will be reached at a pH close to 7, ✓ which is in the colour change range of bromothymol blue.		2
3	3.1	An acid is a proton donor. ✓	1
	3.2	A concentrated acid is an acid solution with a high proportion of acid (solute) to water (solvent). ✓✓ regardless of the degree of ionisation or dissociation of the acid in the solution. A strong acid is an acid that dissociates or ionises almost completely when in solution, regardless of its concentration. ✓✓	4
4	4.1	LiOH ✓	5
	4.2	H ₂ PO ₄ ✓	
	4.3	HNO ₃ ✓	
	4.4	HCOOH ✓	
	4.5	NaHCO ₃ ✓	
5	$\text{H}_2\text{SO}_4 + 2\text{H}_2\text{O} \checkmark \rightarrow 2\text{H}_3\text{O}^+ \checkmark + \text{SO}_4^{2-} \checkmark$ OR: $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$		3
6	6.1	A standard solution is a solution of known concentration. ✓✓	2
	6.2	$c = \frac{m}{MV} \quad \therefore m = cMV \checkmark$ $= (0,15)(106)(0,25) \checkmark$ $= \underline{3,975 \text{ g Na}_2\text{CO}_3} \checkmark$	3
	6.3	$\frac{n_a}{n_b} = \frac{c_a V_a}{c_b V_b} \quad \therefore c_a = \frac{n_a c_b V_b}{n_b V_a} \checkmark$ $= \frac{(1)(0,15)(25)}{(1)(27,5)} \checkmark$ $= \underline{1,38 \times 10^{-1} \text{ mol.dm}^{-3}} \checkmark$	4
7	pH = -log[H ⁺] ✓ and: HCl dissociates completely $\therefore 0,2 \text{ mol.dm}^{-3} \text{ H}^+ \text{ ions}$ = -log[0,2] ✓ = 0,69 ✓		3

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Lingjun Ying



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molecular orbital MO theory of bonding galvanic cells and chemical thermodynamics are considered next Separate chapters are devoted to MO theory of covalent and metallic bonding orbital hybridization intermolecular forces acids and bases ionic equilibrium calculations and polymers and biochemicals This monograph is intended for chemistry students Molecular World Catherine M. Jackson, 2023-07-04 A compelling and innovative account that reshapes our view of nineteenth century chemistry explaining a critical period in chemistry's quest to understand and manipulate organic nature According to existing histories theory drove chemistry's remarkable nineteenth century development In Molecular World Catherine M Jackson shows instead how novel experimental approaches combined with what she calls laboratory reasoning enabled chemists to bridge wet chemistry and abstract concepts and in so doing create the molecular world Jackson introduces a series of practice based breakthroughs that include chemistry's move into lampworked glassware the field's turn to synthesis and subsequent struggles to characterize and differentiate the products of synthesis and the gradual development of institutional chemical laboratories an advance accelerated by synthesis and the dangers it introduced Jackson's historical reassessment emerges from the investigation of alkaloids by German chemists Justus Liebig August Wilhelm Hofmann and Albert Ladenburg Stymied in his own research Liebig steered his student Hofmann into pioneering synthesis as a new investigative method Hofmann's practice based laboratory reasoning produced a major theoretical advance but he failed to make alkaloids That landmark fell to Ladenburg who turned to cutting edge theory only after his successful synthesis In telling the story of these scientists and their peers Jackson reveals organic synthesis as the ground chemists stood upon to forge a new relationship between experiment and theory with far reaching consequences for chemistry as a discipline **Study and Problem Solving Guide to Accompany Principles of Modern Chemistry, Oxtoby/Nachtrieb** Wade A. Freeman, 1986

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Table of Contents Holt Chemistry Acid And Bases Test Answers

1. Understanding the eBook Holt Chemistry Acid And Bases Test Answers
 - The Rise of Digital Reading Holt Chemistry Acid And Bases Test Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Holt Chemistry Acid And Bases Test Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Holt Chemistry Acid And Bases Test Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Holt Chemistry Acid And Bases Test Answers
 - Personalized Recommendations
 - Holt Chemistry Acid And Bases Test Answers User Reviews and Ratings
 - Holt Chemistry Acid And Bases Test Answers and Bestseller Lists

5. Accessing Holt Chemistry Acid And Bases Test Answers Free and Paid eBooks
 - Holt Chemistry Acid And Bases Test Answers Public Domain eBooks
 - Holt Chemistry Acid And Bases Test Answers eBook Subscription Services
 - Holt Chemistry Acid And Bases Test Answers Budget-Friendly Options
6. Navigating Holt Chemistry Acid And Bases Test Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Holt Chemistry Acid And Bases Test Answers Compatibility with Devices
 - Holt Chemistry Acid And Bases Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Holt Chemistry Acid And Bases Test Answers
 - Highlighting and Note-Taking Holt Chemistry Acid And Bases Test Answers
 - Interactive Elements Holt Chemistry Acid And Bases Test Answers
8. Staying Engaged with Holt Chemistry Acid And Bases Test Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Holt Chemistry Acid And Bases Test Answers
9. Balancing eBooks and Physical Books Holt Chemistry Acid And Bases Test Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Holt Chemistry Acid And Bases Test Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Holt Chemistry Acid And Bases Test Answers
 - Setting Reading Goals Holt Chemistry Acid And Bases Test Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Holt Chemistry Acid And Bases Test Answers
 - Fact-Checking eBook Content of Holt Chemistry Acid And Bases Test Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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