

*Thermodynamics*

**MULTIPLE CHOICE** On the line at the left of each statement, write the letter of the choice that best completes the statement or answers the question.

1. \_\_\_\_\_ If an ideal gas does positive work on its surroundings,
  - a. the temperature of the gas increases.
  - b. the volume of the gas increases.
  - c. the pressure of the gas increases.
  - d. the internal energy of the gas decreases.
2. \_\_\_\_\_ Sunlight strikes an ice cube at its melting point and causes it to melt. What is this process?
  - a. adiabatic process
  - b. isothermal process
  - c. isobaric process
  - d. isovolumetric process
3. \_\_\_\_\_ Which of the following processes for an ideal gas system has an unchanging internal energy and a heat intake that corresponds to the value of the work done by the system?
  - a. isothermal process
  - b. isobaric process
  - c. adiabatic process
  - d. isovolumetric process
4. \_\_\_\_\_ How is conservation of internal energy expressed for an isovolumetric system?
  - a.  $Q = W = \Delta U = 0$ ;  $U_i = U_f$
  - b.  $Q = 0$ ; so  $\Delta U = -W$
  - c.  $\Delta T = 0$ , so  $\Delta U = Q - W = 0$ , so  $Q = W$
  - d.  $\Delta V = 0$ , so  $P\Delta V = 0$  and  $W = 0$ , so  $\Delta U = Q$
5. \_\_\_\_\_ A 4 mol ideal gas system undergoes an adiabatic process in which it expands and does 20 J of work on its environment. What is its change in internal energy?
  - a. -20 J
  - b. -5 J
  - c. 0 J
  - d. 20 J
6. \_\_\_\_\_ According to the second law of thermodynamics, the heat received by a heat engine operating in a complete cycle from a high-temperature reservoir
  - a. must be completely converted to work.
  - b. equals the entropy increase.
  - c. can be completely converted to internal energy.
  - d. cannot be completely converted to work.

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**Mike Jess**



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