

## Chemistry of Life

# Study Guide B

## Answer Key

### SECTION 1. ATOMS, IONS, AND MOLECULES

1. An element is a certain type of atom.
2. Sketch should resemble one of the illustrations in Figure 1.2 in Section 1. Nucleus should be in the center with protons and neutrons. Electrons should be labeled in a ring surrounding the nucleus.
3. A compound is composed of atoms of different elements.
4. an atom that has gained or lost one or more electrons
5. the number of protons no longer equals the number of electrons
6. Sketch should resemble the first illustration in Figure 1.3, with one atom losing an electron (becoming a positive ion) and the other accepting that electron (becoming a negative ion).
7. A bond that forms between two atoms that share a pair of electrons.
8. the number of electron pairs that it needs to share to fill its outer energy level
9. molecule
10. compound
11. element
12. ion
13. Ionic bonds form between oppositely-charged ions; covalent bonds form when atoms share a pair of electrons.

### SECTION 2. PROPERTIES OF WATER

1. a molecule with a slightly positively charged region and a slightly negatively charged region
2. The oxygen nucleus pulls the electrons in the molecule more strongly than the hydrogen atoms. The hydrogen atoms have a slight positive charge and the oxygen atom has a slight negative charge.
3. an attraction between a slightly positive hydrogen and a slightly negative atom
4. between the slightly positive hydrogen atom of one water molecule and the

slightly negative oxygen atom of another water molecule

5. A large amount of energy is needed to produce an increase in water temperature; water resists changes in temperature.
6. Water molecules "stick" to each other.
7. Water molecules "stick" to other substances.
8. A solvent is present in greater concentration and dissolves the solute.
9. ions and polar molecules
10. nonpolar molecules, such as fats and oils
11. **Effect on  $H^+$  concentration:** acids increase  $H^+$  concentration; bases decrease  $H^+$  concentration  
**Effect on pH:** acids lower pH; bases raise pH
12. Sketch should resemble Visual Vocab in Section 2 of the text. Solution: homogeneous mixture of substances; solvent: substance present in greater amount in which other substances dissolve; solute: substance that dissolves in a solvent.

### SECTION 3. CARBON-BASED MOLECULES

1. Carbon atoms are the basis of the molecules that make up most living things.
2. Carbon atoms can form covalent bonds with up to four other atoms, including other carbon atoms.
3. Sketches should resemble Figure 3.1 in Section 3.
4. broken down as a source of chemical energy; part of cell structure
5. sugars, starches, cellulose
6. broken down as a source of chemical energy; part of cell structure
7. fats, oils, phospholipids
8. many functions, including movement, transport, chemical catalysts
9. enzymes, hemoglobin
10. store genetic information, build proteins

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**Wolfgang Guggemos**



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