CHEMISTRY ADVANCED STUDY

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FUNDAMENTAL CONCEPTS

Math

- · Properties of exponents
- · Properties of logarithms

- a log, (M) = p log, (M)
- · Graphing slope integrals
 - o y=mx+b
 - b is the y-intercept, m is the slope
- Quadratic formula:

$$0 \quad X \equiv \frac{-b \pm cb^2 - 4ac}{2a}$$

$$0 \quad x = \frac{-4 \pm \sqrt{14 - 4 + (-21)}}{2}$$

$$6 x = -2 \pm 5$$

Science

- Scientific notation
- Dimensional analysis
- Significant figures

Physics and Biology

- Scientific method.
- Coulomb's law: electric force between like or opposite charges; dependence on separation distance and charge magnitude

ELEMENTS AND ATOMS

Protons: define the element (atomic number), positive

Electrons: define the charge (bonds and polarity), negative

Neutrons; define isotopes, neutral

ATOMIC MASS AND WEIGHT

Atomic Mass: Carbon-12 (exactly 12 atomic mass unit)

Atomic Weight

- · weighted average of atomic masses of element
- (atomic mass neutrons) + (atomic mass neutrons)

MOLE A.K.A. AVOGADRO'S NUMBER

× 6.02 - 10P

In Terms of Carbon

1 mole of "C atoms in 12g of Carbon 6.02 - 10²³ = ¹²C atoms in 12g of Carbon

Gram to AMU

1 gram = 1 mole of atomic mass units.

1-gram = 6.02 - 10²² amu

1 mai PAI = 27g

2 mail 253 + 56a

4 moi 110 = 64g

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C Cleary

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