

Lab Report: Gravimetric Analysis of an Unknown Sulfate

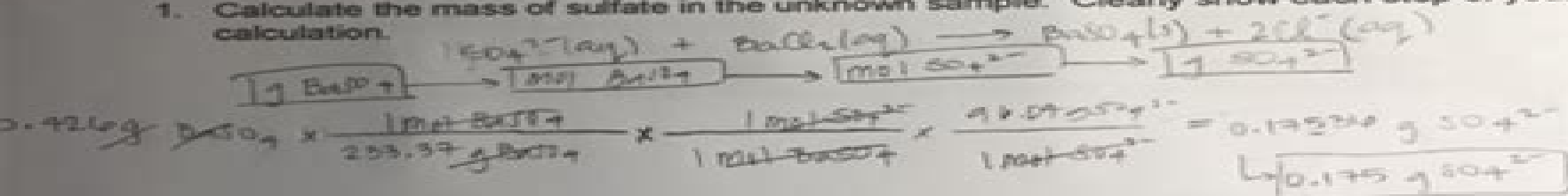
Experimental Data

Unknown Sulfate ID Code: 20

1. Mass of empty 250-mL beaker	111.244 g
2. Mass of 250-mL beaker and unknown sulfate	112.227 g
3. Mass of unknown sulfate	0.343 g
4. Mass of empty crucible (without lid)	11.203 g
5. Mass of crucible (without lid) and barium sulfate	11.629 g
6. Mass of barium sulfate	0.426 g

Calculations and Conclusions

- Calculate the mass of sulfate in the unknown sample. Clearly show each step of your calculation.



- Calculate the percentage by mass of sulfate in the unknown sample. Show your work.
mass of unknown sulfate: 0.343 g ; mass of sulfate: 0.175 g SO_4^{2-}

$$\% \text{ SO}_4^{2-} = \frac{0.175 \text{ g}}{0.343 \text{ g}} \times 100 = 51.0204\%$$

$$\hookrightarrow \boxed{51.0\%}$$

- Calculate the percentage by mass of the metal in the unknown sample. Show your work.

$$100.0\% - 51.0\% = 49.0\%$$

Gravimetric Analysis Lab Report Ucsd

JA Banks



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Quantitative Chemical Analysis Daniel C. Harris, 2010-04-30 QCA is the bestselling textbook of choice for analytical chemistry. It offers a modern portrait of the techniques of chemical analysis backed by a wealth of real world applications. This edition features new coverage of spectroscopy and statistics, new pedagogy and enhanced lecturer support. *Nuclear Science Abstracts*, 1975-08 *Energy Research Abstracts*, 1985 **Scientific and Technical Aerospace Reports**, 1971

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