

Name: \_\_\_\_\_

## Industrial Revolution Study Guide

Fill in the following statements:

1. What is laissez-faire economics?
2. What is capitalism?
3. What is a factory?
4. What is a stock?
5. What is a corporation?
6. Name five inventions that came from the Industrial Revolution.
  - 1.
  - 2.
  - 3.
  - 4.
  - 5.
7. What is a textile?
8. Where AND when did the Industrial Revolution begin?

Where:

When:
9. What are the inventions we discussed of Thomas Edison?
10. What did Alexander Graham Bell invent?
11. What did Marconi invent?
12. What is the assembly line and who invented it?

# Industrial Revolution Guide Answers

**M Walker**



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web limiting reactant theoretical and percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$  g if a reaction vessel contains 0

mol  $\text{CO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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$$4\text{Cu} + 2\text{AgNO}_3 \rightarrow 2\text{Cu} + 2\text{Ag} + 2\text{Cu(NO}_3)_2$$

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web start with 8 3 mole  $\text{H}_2\text{S}$  and multiply it with 2 mole  $\text{H}_2\text{O}$  over 2 mole  $\text{H}_2\text{S}$  from the equation  $8 3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149 4 \text{ g } \text{H}_2\text{O}$  is your theoretical yield so percent yield  $\frac{137 1 \text{ g } \text{H}_2\text{O}}{149 4 \text{ g } \text{H}_2\text{O}} \times 100 = 91 77\%$  problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction  $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$  b determine the theoretical yield of  $\text{Al}_2\text{O}_3$  160.0 g  $\text{Al}$   $\times \frac{1 \text{ mol Al}}{27.0 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$  c determine the percent yield  $\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$

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web 7  $\text{Na}_2\text{CO}_4$  is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12  $\text{C}$  is the limiting reagent 4 33 g of  $\text{H}_2$  are left over 13

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 1 333.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g  $\text{CO}$  1 mol  $\text{CO}$

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text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3

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