

## QUESTION 5

- 5.1
- 5.1.1 Aufbau diagram ✓✓ (2)
  - 5.1.2 Sulphur ✓✓ (2)
  - 5.1.3  $1s^2 2s^2 2p^4 3s^2 3p^2$  ✓✓ (2)
  - 5.1.4 Group 16 ✓  
Period 3 ✓ (2)
  - 5.1.5 The element has 6 valence electrons ✓ and belong to the third horizontal row of the periodic table. ✓ (2)
  - 5.1.5 Non-metal ✓ (1)
  - 5.2.1 5 and 1, respectively (1)
  - 5.2.2  $\begin{array}{c} \text{Y} \text{ : } \text{X} \text{ : } \text{Y} \\ \vdots \\ \text{Y} \end{array}$  ✓ (2)
  - 5.2.4 N ✓ and H ✓, or P and H (2)
- [18]

## QUESTION 6

- 6.1
- (a) 24 ✓
  - (b) 12 ✓
  - (c)  $\text{Na}^+$  ✓
  - (d) 10 ✓
  - (e) 12 ✓
  - (f) 17 ✓
  - (g) 19 ✓
- 6.2 Relative atomic mass =  $(0,76 \times 35)$  ✓ +  $(0,24 \times 37)$  ✓  
= 35,48 ✓ (3)
- 6.3 Isotopes ✓✓ (2)
- [12]

## QUESTION 7

- 7.1 Sodium hydroxide ( $\text{NaOH}$ ) ✓  
Hydrogen ✓ or  $\text{H}_2(\text{g})$  (2)
- 7.2  $2\text{Mg}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{MgO}(\text{s})$  ✓✓ (bal) (3)
- 7.3 Sodium + water → sodium hydroxide + water (5)
- $2\text{Na}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \rightarrow 2\text{NaOH}(\text{aq}) + \text{H}_2(\text{g})$  ✓ ✓ (Bal)
- 7.4 Because they react vigorously with air to form oxides ✓ and with water to form hydroxides. ✓ (2)
- 7.5 It is an inert colourless gas ✓  
It is unreactive ✓ (2)
- 7.6 Noble gases (1)
- 7.7 Because they don't react, it was not easy to discover them, so no one knew they existed. ✓✓ (2)

# Grade11 June Exemplar 2014

**SB Merriam**



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### **Table of Contents Grade11 June Exemplar 2014**

1. Understanding the eBook Grade11 June Exemplar 2014
  - The Rise of Digital Reading Grade11 June Exemplar 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade11 June Exemplar 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade11 June Exemplar 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade11 June Exemplar 2014
  - Personalized Recommendations
  - Grade11 June Exemplar 2014 User Reviews and Ratings
  - Grade11 June Exemplar 2014 and Bestseller Lists

5. Accessing Grade11 June Exemplar 2014 Free and Paid eBooks
  - Grade11 June Exemplar 2014 Public Domain eBooks
  - Grade11 June Exemplar 2014 eBook Subscription Services
  - Grade11 June Exemplar 2014 Budget-Friendly Options
6. Navigating Grade11 June Exemplar 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Grade11 June Exemplar 2014 Compatibility with Devices
  - Grade11 June Exemplar 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade11 June Exemplar 2014
  - Highlighting and Note-Taking Grade11 June Exemplar 2014
  - Interactive Elements Grade11 June Exemplar 2014
8. Staying Engaged with Grade11 June Exemplar 2014
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade11 June Exemplar 2014
9. Balancing eBooks and Physical Books Grade11 June Exemplar 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade11 June Exemplar 2014
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Grade11 June Exemplar 2014
  - Setting Reading Goals Grade11 June Exemplar 2014
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade11 June Exemplar 2014
  - Fact-Checking eBook Content of Grade11 June Exemplar 2014
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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