

ACID-BASE STUDY GUIDE

- Normal Plasma pH 7.35 to 7.45 (hydrogen (H⁺) concentration)
 - Higher this goes, higher the mortality rate
 - Measure the acidity or alkalinity of blood
 - Homeostatic mechanisms keep pH normal
 - Buffer systems, kidneys & lungs
 - Greater H⁺ concentration = lower pH & more acidic solute
 - Lower H⁺ concentration = higher pH & more alkaline solute
- Major ECF buffer system
 - Bicarbonate-carbonic acid system
 - $\text{CO}_2 + \text{H}_2\text{O} \leftrightarrow \text{H}_2\text{CO}_3 \leftrightarrow \text{H}^+ \text{ and } \text{HCO}_3^-$
 - Kidneys regulate bicarbonate in the ECF
 - Lungs (controlled by medulla) regulate CO₂, thus the carbonic acid in the ECF
 - Lungs are first responders
- Other buffer systems
 - ECF: inorganic phosphates and plasma proteins
 - ICF: proteins, organic, inorganic phosphates
 - Hemoglobin

METABOLIC ACIDOSIS

pH lower than 7.35

- Can occur by a gain of H⁺ or loss of bicarbonate ions (<22 mm Hg)
- Causes
 - Direct loss of bicarbonate (diarrhea, administration of IV nutrition w/o bicarb or bicarb-producing solutes, diuretic use, early renal insufficiency, excess chloride, lower intestinal fistulas, ureterostomies)
 - Most common due to kidney injury
 - Lactic acidosis (higher it is, the worse sepsis is), salicylate poisoning, renal failure, methanol, ethylene, or propylene toxicity, DKA, ketoacidosis due to starvation
 - Systemic issue: diabetes
- Clinical Manifestations
 - Headache, confusion, drowsiness, increased RR/depth, low BP, nausea, vomiting, cold/clammy skin, arrhythmias, shock, peripheral vasodilation, decreased cardiac output
- Medical Management
 - First have to correct underlying issue

Key Acid Base Study Guide

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Gloffke, George Odian**



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