

UNIT (6) ACIDS AND BASES

6.1 | Arrhenius Definition of Acids and Bases

Definitions for acids and bases were proposed by the Swedish chemist *Savante Arrhenius* in 1884.

Acids were defined as compounds that contain hydrogen and can dissolve in water to release hydrogen ions into solution. For example, HCl dissolves in water as follows:



Bases are defined as substances that dissolve in water to release hydroxide ions (OH^-) into solution. For example, a typical base according to the Arrhenius definition is sodium hydroxide, NaOH:



The process that converts a molecule such as HCl into ions is called ionization.

Ionization is the production of ions from molecular compounds. No ions were initially present.

The process of converting a solid ionic compound, such as NaOH, into ions in solution is called dissociation.

Dissociation is the production of ions from ionic compounds. Ions were initially present, water just breaks apart ions.

Some common acidic and basic substances:

Acidic	Basic
citrus fruits	soap
vinegar	detergents
aspirin	milk of magnesia

Some characteristic properties associated with acids and bases:

Acids	Bases
sour taste	bitter taste
change litmus paper to red	change litmus paper to blue
react with hydroxide bases to produce salt	slippery, soapy feeling

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Table of Contents Harcourtschoolsupply Answers Unit 1

1. Understanding the eBook Harcourtschoolsupply Answers Unit 1
 - The Rise of Digital Reading Harcourtschoolsupply Answers Unit 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Harcourtschoolsupply Answers Unit 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Harcourtschoolsupply Answers Unit 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Harcourtschoolsupply Answers Unit 1
 - Personalized Recommendations
 - Harcourtschoolsupply Answers Unit 1 User Reviews and Ratings
 - Harcourtschoolsupply Answers Unit 1 and Bestseller Lists

5. Accessing Harcourtschoolsupply Answers Unit 1 Free and Paid eBooks
 - Harcourtschoolsupply Answers Unit 1 Public Domain eBooks
 - Harcourtschoolsupply Answers Unit 1 eBook Subscription Services
 - Harcourtschoolsupply Answers Unit 1 Budget-Friendly Options
6. Navigating Harcourtschoolsupply Answers Unit 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Harcourtschoolsupply Answers Unit 1 Compatibility with Devices
 - Harcourtschoolsupply Answers Unit 1 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Harcourtschoolsupply Answers Unit 1
 - Highlighting and Note-Taking Harcourtschoolsupply Answers Unit 1
 - Interactive Elements Harcourtschoolsupply Answers Unit 1
8. Staying Engaged with Harcourtschoolsupply Answers Unit 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Harcourtschoolsupply Answers Unit 1
9. Balancing eBooks and Physical Books Harcourtschoolsupply Answers Unit 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Harcourtschoolsupply Answers Unit 1
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Harcourtschoolsupply Answers Unit 1
 - Setting Reading Goals Harcourtschoolsupply Answers Unit 1
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Harcourtschoolsupply Answers Unit 1
 - Fact-Checking eBook Content of Harcourtschoolsupply Answers Unit 1
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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