

2

MATTER AND CHANGE

SECTION 2.1 PROPERTIES OF MATTER (pages 39–42)

This section helps you distinguish extensive from intensive properties and identify substances by their properties. It teaches you how to differentiate the three states of matter. It also defines a physical property and lists examples of physical properties and physical changes.

► Describing Matter (page 39)

1. The _____ mass _____ of an object is a measure of the amount of matter the object contains.

2. How does an extensive property differ from an intensive property?

An extensive property depends on the amount of matter, while an intensive property depends on the type of matter.

► Identifying Substances (page 40)

3. Matter that has a uniform and definite composition is called a _____ substance _____.

4. Is the following sentence true or false? All samples of a substance have different physical properties. _____ false _____

5. A physical property is a quality or condition of a substance that can be _____ observed _____ or _____ measured _____ without changing the substance's composition.

6. Circle the letter of the term that is NOT a physical property.

- a. hardness c. boiling point
b. color **d. melting**

7. Look at Table 2.1 on page 40. What is the melting point of bromine? _____ -7°C _____

8. Look at Table 2.1 on page 40. Circle the letter of the substance that is a yellow solid and melts at 115°C.

- a. sulfur**
b. chlorine
c. gold
d. copper

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Yicheng Fang



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Table of Contents Guided Reading And Study Workbook Chapter 2 Matter Change

1. Understanding the eBook Guided Reading And Study Workbook Chapter 2 Matter Change
 - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 2 Matter Change
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 2 Matter Change
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Reading And Study Workbook Chapter 2 Matter Change
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 2 Matter Change

- Personalized Recommendations
- Guided Reading And Study Workbook Chapter 2 Matter Change User Reviews and Ratings
- Guided Reading And Study Workbook Chapter 2 Matter Change and Bestseller Lists
- 5. Accessing Guided Reading And Study Workbook Chapter 2 Matter Change Free and Paid eBooks
 - Guided Reading And Study Workbook Chapter 2 Matter Change Public Domain eBooks
 - Guided Reading And Study Workbook Chapter 2 Matter Change eBook Subscription Services
 - Guided Reading And Study Workbook Chapter 2 Matter Change Budget-Friendly Options
- 6. Navigating Guided Reading And Study Workbook Chapter 2 Matter Change eBook Formats
 - ePub, PDF, MOBI, and More
 - Guided Reading And Study Workbook Chapter 2 Matter Change Compatibility with Devices
 - Guided Reading And Study Workbook Chapter 2 Matter Change Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 2 Matter Change
 - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 2 Matter Change
 - Interactive Elements Guided Reading And Study Workbook Chapter 2 Matter Change
- 8. Staying Engaged with Guided Reading And Study Workbook Chapter 2 Matter Change
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Reading And Study Workbook Chapter 2 Matter Change
- 9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 2 Matter Change
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 2 Matter Change
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 2 Matter Change
 - Setting Reading Goals Guided Reading And Study Workbook Chapter 2 Matter Change
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 2 Matter Change

- Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 2 Matter Change
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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