Acids and bases

Chapter outline

- Recall the definitions of Bronsted-Lowry acids and bases and deduce these in a chemical reaction, including the identification of conjugate acid-base pairs.
- · Recall the terms 'amphoteric' and 'amphiprotic'.
- Describe the reactions of acids with metals, metal oxides, metal hydroxides, hydrogen carbonates and carbonates.
- Recall that pH is a measure of the concentration of aqueous hydrogen ions and solve problems involving pH and the concentration of hydrogen ions and hydroxide ions.
- Explain the difference between strong and weak acids and bases and describe experiments that could be used to distinguish these.
- Recall what is meant by acid deposition and explain some of the problems associated with acid deposition as well as pre- and post-combustion methods of reducing sulfur oxide emissions.
- Recall the definitions of Lewis acids and bases and recognise these in a chemical reaction.
- Deduce the dissociation constant for weak acids and bases.
- Sketch and explain the important features of titration curves including the composition and action of a buffer and use these curves to deduce the pK₃ or pK₆, value of a weak acid or base.
- Explain how an acid-base indicator works and select a suitable indicator for a titration.

KEY TERMS

Acid deposition: A more general term than acid rain; it refers to any process in which acidic substances leave the atmosphere and are deposited on the Earth's surface.

Brensted-Lowry acid: A proton (H*) donor.

Bronsted-Lowry base: A proton (H*) acceptor.

pH: A measure of the concentration of hydrogen ions in an aqueous solution, $pH = -log_{10}[H^*]$

Strong acid/base: An acid/base that is completely dissociated into its ions in aqueous solution.

Weak acid/base: An acid/base that is only partially dissociated into its ions in aqueous solution.

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M Walker

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