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web a balanced chemical equation often may be derived from a qualitative description of some chemical reaction by a fairly simple approach known as balancing by inspection consider as an example the decomposition of water to yield molecular hydrogen and oxygen

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web chemical equations must be balanced so that the quantities of reactants and products match for an equation to be balanced there must be an equal numbers of atoms on each side part of chemistry

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web oct 31 2023 to balance a chemical equation enter an equation of a chemical reaction and press the balance button the balanced equation will appear above use uppercase for the first character in the element and lowercase for the second character examples fe au co br c o n f ionic charges are not yet supported and will be ignored

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web feb 19 2022 in a balanced chemical equation both the numbers of each type of atom and the total charge are the same on both sides equations ref 3 1 1 and ref 3 1 2 are balanced chemical equations what is different on each side of the equation is how the atoms are arranged to make molecules or ions

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web mar 14 2023 the ultimate goal for balancing chemical equations is to make both sides of the reaction the reactants and the products equal in the number of atoms per element this stems from the universal law of the conservation of mass which states that matter can neither be created nor destroyed

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web balancing equations the chemical equation described in section 4 1 is balanced meaning that equal numbers of atoms for each element involved in the reaction are represented on the reactant and product sides this is a requirement the equation

must satisfy to be consistent with the law of conservation of matter

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web because of the law of conservation of mass balancing chemical equations gives a more accurate representation on what s happening in a reaction the accuracy of the ratios of moles is important in calculating the theoretical yield and how much stuff is reacting mass wise stoichiometry

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web a key step in balancing chemical equations is correctly identifying the formulas of the reactants and products for example consider the reaction between calcium oxide CaO s and H_2O l to form aqueous calcium hydroxide b is it possible to balance the equation if you incorrectly identify the product as CaOH aq2 and if so what is the equation

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web dec 2 2020 a balanced chemical equation tells you the amounts of reactants and products needed to satisfy the law of conservation of mass basically this means there are the same numbers of each type of atoms on the left side of the equation as there are on the right side of the equation

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web balance the following chemical equation $\text{Mg} + \text{OH}^- + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2\text{O}$ note all reactants and products require a coefficient of at least one stuck review related articles videos or use a hint report a problem do 4 problems

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web chemists use symbols and formulae to represent elements and compounds word equations and balanced chemical equations represent the changes that happen in chemical reactions

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web oct 13 2023 method 1 doing a traditional balance download article 1 write down your given equation for this example you will use $\text{C}_3\text{H}_8 + 2 \text{H}_2\text{O} \rightarrow \text{CO}_2$ this reaction occurs when propane C_3H_8 is burned in the presence of oxygen to produce

water and carbon dioxide 2 write down the number of atoms per element

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web the ions which do not participate in chemical reactions and present the same on both sides of the reactions are called spectator ions in the net chemical reaction the spectator ions are cancelled from both sides of the equation in other words spectate means watch the other ions react in an aqueous solution

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web answer $\text{Na}^+ \text{aq}$ and $\text{SO}_4^{2-} \text{aq}$ identify the chemical equation the spectator ions $\text{KOH} \text{aq}$ HNO_3 $\text{KNO}_3 \text{aq}$ H_2O $\text{K}^+ \text{aq}$ $\text{OH}^- \text{aq}$ $\text{H}^+ \text{aq}$ $\text{NO}_3^- \text{aq}$ $\text{K}^+ \text{aq}$ $\text{NO}_3^- \text{aq}$ H_2O $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ identify the chemical equation the spectator ions $\text{Pb}^{2+} \text{aq}$ $2 \text{KI} \text{aq}$ $\text{PbI}_2 \text{s}$ $2 \text{KNO}_3 \text{aq}$

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web these ions are called spectator ions because they do not participate in the actual reaction canceling the spectator ions gives the net ionic equation which shows only those species that participate in the chemical reaction $2 \text{Ag}^+ \text{aq}$ $\text{Cr}_2\text{O}_7^{2-} \text{aq}$ H_2O $2 \text{Cr}^{3+} \text{aq}$ $2 \text{OH}^- \text{aq}$

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web chemical equation in which all dissolved ionic reactants and products including spectator ions are explicitly represented by formulas for their dissociated ions molecular equation chemical equation in which all reactants and products are

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web jan 29 2020 the ionic form of this reaction is $2 \text{Na}^+ \text{aq}$ $2 \text{Cl}^- \text{aq}$ $\text{Cu}^{2+} \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $2 \text{Na}^+ \text{aq}$ $\text{SO}_4^{2-} \text{aq}$ $\text{CuCl}_2 \text{s}$ the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six $\text{NO}_3^- \text{aq}$ ions and the six $\text{Na}^+ \text{aq}$ ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation $3 \text{Ba}^{2+} \text{aq}$ $2 \text{PO}_4^{3-} \text{aq}$ $\text{Ba}_3(\text{PO}_4)_2 \text{s}$

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web these spectator ions ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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 $\text{AgNO}_3(\text{aq}) + \text{KCl}(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{KNO}_3(\text{aq})$ total ionic $\text{Ag}^+(\text{aq}) + \text{NO}_3^-(\text{aq}) + \text{K}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s}) + \text{K}^+(\text{aq}) + \text{NO}_3^-(\text{aq})$ net ionic $\text{Ag}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{AgCl}(\text{s})$

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ the spectator ions in $\text{KOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$ are $\text{K}^+(\text{aq})$ and $\text{NO}_3^-(\text{aq})$ how to identify and list the

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web aug 8 2022 a spectator ion is an ion that does not take part in the chemical reaction and is found in solution both before and after the reaction the net ionic equation is the chemical equation that shows only those elements compounds and ions that are directly involved in the chemical reaction

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