

Ileap Interpretive Guide

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Ileap Interpretive Guide

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a net ionic equation is the most accurate representation of the actual chemical process that occurs writing these equations requires a familiarity with solubility rules acid base reactivity

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is $2\text{Cl}^- \text{aq} + 2\text{Ag}^+ \text{aq}$

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver ions regardless of the source of these ions these molecular and

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the complete ionic equation for this reaction is as follows equation 4 $13 \text{ Ag}^+ \text{ aq} + 2 \text{ F}^- \text{ aq} + 2 \text{ NH}_4^+ \text{ aq} + \text{Cr}_2\text{O}_7^{2-} \text{ aq} \rightarrow \text{Ag}_2\text{Cr}_2\text{O}_7 \text{ s} + 2 \text{ NH}_4^+ \text{ aq} + 2 \text{ F}^- \text{ aq}$ because two $\text{NH}_4^+ \text{ aq}$

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