

## **EXPERIMENT NO.1** **EXPERIMENT NAME: Redox Titration**

**Objective:** - Find out the percentage of iron in a given iron alloy solution, when X grams of alloy dissolved in 1 liter

**Apparatus/Reagent Required :-** Burette, Pipette, Conical flask, Funnel,  
Std  $\text{KMnO}_4$  solution,  $\text{KMnO}_4$  solution, dil  $\text{H}_2\text{SO}_4$ , Alloy solution.

**Theory:** -

To determine the amount of iron in alloy, the alloy sample is dissolved in dil.  $\text{H}_2\text{SO}_4$ , the iron present is converted to ferrous sulfate with the liberation of hydrogen gas.



The  $\text{Fe}^{2+}$  present in the solution can be determined by titration with a standard  $\text{KMnO}_4$  solution. Potassium permanganate oxidizes ferrous ion into ferric ion in acid medium in cold and is itself reduced to colourless manganous ion.



The appearance of a faint but permanent pink colour marks the end point.

**Chemical Equation:-**



**Procedure:** -

- Fill burette with  $\text{KMnO}_4$  solution.

# Lorven Publications Engg Chemistry Lab Manuals

**Roman Wölfel**



**Lorven Publications Engg Chemistry Lab Manuals:**

## Whispering the Strategies of Language: An Mental Journey through **Lorven Publications Engg Chemistry Lab Manuals**

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