

## 12

## STOICHIOMETRY

**SECTION 12.1 THE ARITHMETIC OF EQUATIONS** (pages 353–358)

*This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.*

**► Using Everyday Equations** (pages 353–355)

1. How can you determine the quantities of reactants and products in a chemical reaction?

You can use the balanced equation.

2. Quantity usually means the amount of a substance expressed in grams or moles.

3. A bookcase is to be built from 3 shelves (Sh), 2 side boards (Sb), 1 top (T), 1 base (B), and 4 legs (L). Write a “balanced equation” for the construction of this bookcase.

$3\text{Sh} + 2\text{Sb} + \text{T} + \text{B} + 4\text{L} = \text{Sh}_3\text{Sb}_2\text{TBL}_4$

**► Using Balanced Chemical Equations** (page 354)

4. Is the following sentence true or false? Stoichiometry is the calculation of quantities in chemical reactions. true

5. Calculations using balanced equations are called stoichiometric calculations.

**► Interpreting Chemical Equations** (pages 356–357)

6. From what elements is ammonia produced? How is it used?

Ammonia molecules are composed of nitrogen and hydrogen; it is used as a fertilizer.

7. Circle the letter of the term that tells what kind of information you CANNOT get from a chemical equation.

a. moles

d. volume

b. mass

e. number of particles

☒ c. size of particles

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**M Lipman**



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## **Table of Contents Guided And Study Workbook Practice Problem**

1. Understanding the eBook Guided And Study Workbook Practice Problem
  - The Rise of Digital Reading Guided And Study Workbook Practice Problem
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided And Study Workbook Practice Problem
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided And Study Workbook Practice Problem
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided And Study Workbook Practice Problem

- Personalized Recommendations
  - Guided And Study Workbook Practice Problem User Reviews and Ratings
  - Guided And Study Workbook Practice Problem and Bestseller Lists
5. Accessing Guided And Study Workbook Practice Problem Free and Paid eBooks
    - Guided And Study Workbook Practice Problem Public Domain eBooks
    - Guided And Study Workbook Practice Problem eBook Subscription Services
    - Guided And Study Workbook Practice Problem Budget-Friendly Options
  6. Navigating Guided And Study Workbook Practice Problem eBook Formats
    - ePub, PDF, MOBI, and More
    - Guided And Study Workbook Practice Problem Compatibility with Devices
    - Guided And Study Workbook Practice Problem Enhanced eBook Features
  7. Enhancing Your Reading Experience
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    - Highlighting and Note-Taking Guided And Study Workbook Practice Problem
    - Interactive Elements Guided And Study Workbook Practice Problem
  8. Staying Engaged with Guided And Study Workbook Practice Problem
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Guided And Study Workbook Practice Problem
  9. Balancing eBooks and Physical Books Guided And Study Workbook Practice Problem
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Guided And Study Workbook Practice Problem
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Guided And Study Workbook Practice Problem
    - Setting Reading Goals Guided And Study Workbook Practice Problem
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Guided And Study Workbook Practice Problem

- Fact-Checking eBook Content of Guided And Study Workbook Practice Problem
  - Distinguishing Credible Sources
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- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
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