

Isotopes Practice Set

- What does atomic number tell you about an element?
- What does the number used to isotopes signify?
- How can you tell isotopes apart?

For each of the following isotopes, write the number of protons, neutrons, and electrons.

	Chlorine-35	Chlorine-37
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Nitrogen-14	Nitrogen-15
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Carbon-12	Carbon-13
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Oxygen-16	Oxygen-18
# of protons	8	8
# of electrons	8	8
# of neutrons	8	10

Fill in the isotopes names including chemical symbol, atomic number, and mass number from the values that your partner gave and the information provided.

	Isotope: $^{35}_{17}\text{Cl}$	Isotope: $^{37}_{17}\text{Cl}$
# of protons	17	17
# of neutrons	18	20
# of electrons	17	17

	Isotope: $^{12}_6\text{C}$	Isotope: $^{13}_6\text{C}$
# of protons	6	6
# of neutrons	6	7
# of electrons	6	6

	Isotope: $^{14}_7\text{N}$	Isotope: $^{15}_7\text{N}$
# of protons	7	7
# of neutrons	7	8
# of electrons	7	7

	Isotope: $^{16}_8\text{O}$	Isotope: $^{18}_8\text{O}$
# of protons	8	8
# of neutrons	8	10
# of electrons	8	8

Isotopes Practice Set Answer Key

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Table of Contents Isotopes Practice Set Answer Key

1. Understanding the eBook Isotopes Practice Set Answer Key
 - The Rise of Digital Reading Isotopes Practice Set Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Isotopes Practice Set Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Isotopes Practice Set Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Isotopes Practice Set Answer Key
 - Personalized Recommendations
 - Isotopes Practice Set Answer Key User Reviews and Ratings
 - Isotopes Practice Set Answer Key and Bestseller Lists
5. Accessing Isotopes Practice Set Answer Key Free and Paid eBooks
 - Isotopes Practice Set Answer Key Public Domain eBooks
 - Isotopes Practice Set Answer Key eBook Subscription Services
 - Isotopes Practice Set Answer Key Budget-Friendly Options

6. Navigating Isotopes Practice Set Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Isotopes Practice Set Answer Key Compatibility with Devices
 - Isotopes Practice Set Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Isotopes Practice Set Answer Key
 - Highlighting and Note-Taking Isotopes Practice Set Answer Key
 - Interactive Elements Isotopes Practice Set Answer Key
8. Staying Engaged with Isotopes Practice Set Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Isotopes Practice Set Answer Key
9. Balancing eBooks and Physical Books Isotopes Practice Set Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Isotopes Practice Set Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Isotopes Practice Set Answer Key
 - Setting Reading Goals Isotopes Practice Set Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Isotopes Practice Set Answer Key
 - Fact-Checking eBook Content of Isotopes Practice Set Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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