

Chapter 3

Fuel, Air, and Combustion Thermodynamics

3.1) What is the molecular weight, enthalpy (kJ/kg), and entropy (kJ/kg K) of a gas mixture at $P = 1000$ kPa and $T = 500$ K, if the mixture contains the following species and mole fractions?

a) A table of the given and computed parameters is:

i	y_i	M_i [kg/kmol]	$h_{f,i}^\circ$ [kJ/kmol]	$h_i - h_{f,i}^\circ$ [kJ/kmol]	s_i° [kJ/kmol-K]	$y_i (s_i^\circ - R_u \ln(y_i))$ [kJ/kmol-K]	
CO ₂	1	0.10	44.01	-393.52	33.40	269.30	28.84
H ₂ O	2	0.15	18.01	-241.83	26.00	232.74	37.28
N ₂	3	0.70	28.01	0	21.46	228.17	161.79
CO	4	0.05	28.01	-110.53	21.69	234.54	12.97

The mixture molecular mass is:

$$M = \sum y_i M_i = (0.10)(44.01) + (0.15)(18.01) + (0.70)(28.01) + (0.05)(28.01)$$

$$M = 28.11 \text{ kg/kmol}$$

The specific mixture enthalpy is:

$$h_f = \sum y_i h_i = (0.10)(-393.52 + 33.40) + (0.15)(-241.83 + 26.00)$$

$$+ (0.70)(0 + 21.46) + (0.05)(-110.53 + 21.69)$$

$$h_f = -57.8 \text{ kJ/kmol} = -5.78 \times 10^4 \text{ kJ/kmol}$$

The mixture enthalpy is:

$$h = \frac{h_f}{m} = \frac{-5.78 \times 10^4}{28.11}$$

$$h = -2056 \text{ kJ/kg}$$

The specific mixture entropy is:

$$s = -R_u \ln \left(\frac{P}{P_0} \right) + y_i (s_i^\circ - R_u \ln(y_i))$$

$$s = (-8.314) \ln \left(\frac{1000}{100} \right) + 28.84 + 37.28 + 161.79 + 12.97$$

$$s = 221.74 \text{ kJ/kmol-K}$$

Internal Combustion Engines Applied Thermosciences Solutions Manual

L Cohen



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