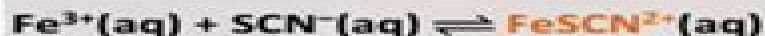


IA lab work on spectrophotometric determination of K_c of the Iron(3)Thiocyanate system



$$K = \frac{[\text{FeSCN}^{2+}]}{[\text{Fe}^{3+}][\text{SCN}^{-}]}$$

Since I will not change the temperature in my experiment, the equilibrium constant will stay the same, and therefore I can measure the light absorbance and use the Lambert Beer-Law to find the initial concentration of each species in the system and then calculate the K_c .

$$X(\text{initial concentration of Iron3 and Thiocyanate}) = A(\text{absorbance of Iron(3)Thiocyanate solution}) / \epsilon \cdot l$$

Equipment:

- 4 test tubes with 4 corks
- test tube rack
- 10 ml pipette with pump thing

Solutions:

- 20 ml of Iron (3) solution
- 20 ml of Thiocyanate solution
- 20 ml of water H_2O

Samples	The amount of Thiocyanate solution SCN^{-}	The amount of water H_2O	Iron 3 solution Fe^{3+}
1	2 ml	8 ml	5 ml
2	4 ml	6 ml	5 ml
3	5 ml	5 ml	5 ml
4	9 ml	1 ml	5 ml

Kc Of Iron Thiocyanate Lab Report

RJ Alexander



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Chemical Synergies Nuno A.G. Bandeira, Bartosz Tylkowski, 2018-05-07 This book gives an overview of recent integrated and inter disciplinary approaches between chemical experiment and theory in a variety of fields from polymer science to materials chemistry and ranging from the design of tailored properties to catalysis and reactivity building on the well established success of Density Functional Theory as the foremost quantum chemical method to provide qualitative and quantitative interpretation of results from the chemical laboratory The combination of several characterization techniques with an understanding at the molecular level of chemical and physical phenomena are the main focal point of the subject matter

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