

12

STOICHIOMETRY

Practice Problems

In your notebook, solve the following problems.

SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate, $2\text{KClO}_3(\text{s})$. Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from 3.8×10^{23} molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



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Pamela Mills, Amina El-Ashmawy



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Essential Concepts of Chemistry Study Guide James R. Braun, Sherman, 1999 *Resources in Education*, 1995

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 $5\text{ 00 l } \text{n}_2 \text{ n}_2\text{ 1 l } 13\text{ l } \text{cl}_2 \times \text{n}_2\text{ 3 l } \text{cl}_2\text{ 4 3 l } \text{n}_2$ answer $4\text{ 3} \times 10^3\text{ ml } \text{n}_2$ b how many grams of chlorine must react to produce
16 l of nitrogen gas at 1 2 atm and 23oc $\text{nn}_2\text{ 1 2atm} \times 16\text{ l}$

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web he determined that if certain gases that are products and reactions in a chemical reaction are measured at the same
conditions temperature and pressure then the volume of gas consumed produced is equal to the ratio between the gases or
the ratio of the coefficients

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web mar 13 2023 solution mark draganjac arkansas state university q1 given the following reaction $\text{h}_2\text{so}_4 + \text{naoh} \rightarrow \text{na}_2\text{so}_4 + \text{h}_2\text{o}$
 $\text{h}_2\text{so}_4 + \text{naoh} \rightarrow \text{na}_2\text{so}_4 + \text{h}_2\text{o}$ if it takes 27 4 ml of 0 768 m naoh to titrate 16 7 ml of h_2so_4 what is the
concentration of the h_2so_4 solution hint balance the equation first q2

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proper units to earn full credit 1 silver and nitric acid react according to the following balanced equation $3 \text{Ag} + 4 \text{HNO}_3 \rightarrow 3 \text{AgNO}_3 + 2 \text{H}_2\text{O} + \text{N}_2\text{O}$

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