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$$2\text{ KClO}_3 \rightarrow 2\text{ KCl} + 3\text{ O}_2$$

$$2\text{ NaCl} + \text{F}_2 \rightarrow 2\text{ NaF} + \text{Cl}_2$$

$$4\text{ H}_2 + \text{O}_2 \rightarrow 2\text{ H}_2\text{O}$$

$$5\text{ Pb} + 2\text{ HCl} \rightarrow 2\text{ H}_2 + \text{PbCl}_2$$

$$2\text{ Al} + 3\text{ K}_2\text{SO}_4 \rightarrow 4\text{ KBr} + \text{Al}_2\text{SO}_4$$

$$3\text{ CH}_4 + 8\text{ O}_2 \rightarrow 3\text{ CO}_2 + 4\text{ H}_2\text{O}$$
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$$1\text{ N}_2 + \text{F}_2 \rightarrow 2\text{ NF}_3$$

$$2\text{ C}_6\text{H}_{10}\text{O}_2 + \text{CO}_2 + \text{H}_2\text{O} \rightarrow 3\text{ HBr} + \text{KHCO}_3 + \text{H}_2\text{O} + \text{KBr} + \text{CO}_2 + \text{GaBr} + \text{NaBr} + \text{SnO}_4 + \text{NaSO}_3 + \text{Ga}_2\text{SO}_3 + 3\text{ N}_2 + 5\text{ NF}_3 + \text{SnF}_2 + \text{N}_2\text{O}_3$$

using the equation from problem 2 above answer the following

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