

Calculations in IB Chemistry SL and HL

Below is a list of the calculations are expected to be able to do in IB chemistry. This needs committing to memory, and each type of calculation will need practice to master. Good luck!

Calculations shaded in grey are only required at HL

Quantitative Chemistry

Purpose	The Calculation	Notes
1. Calculating a quantity in moles from a number of particles	$n = \frac{N}{L}$ n = the quantity in moles N = the number of particles L = Avogadro's constant, 6.02×10^{23}	- When calculating numbers of atoms within molecules, multiply the number of particles (N) by the number of atoms in the formula - To find out numbers of particles, rearrange to: $N = n.L$
2. Determine relative molecular or formula mass, M_r	$M_r = \sum (\text{number of atoms} \cdot \text{atomic mass})$	M_r has no unit as it is a relative value - To calculate molar mass, M_{mol} just stick a 'g' for grams on the end
3. Calculate a quantity in moles, from a mass of a substance	$n = \frac{m}{M_{mol}}$ n = the quantity in moles m = the mass of substance you are given in grams M_{mol} = the molar mass of the substance	- To determine the mass of a given number of moles of a substance use: $m = M_{mol} \cdot n$ - To determine the molar mass of a given mass of a substance: $M_{mol} = \frac{m}{n}$
4. Determine empirical formula from % composition by mass	- Divide each % by the atomic mass - Divide each answer to Step 1 by the smallest answer to Step 1 - Multiply all answers to Step 1 to remove any obvious fractions - If there is a '.5' multiply everything by 2 - If there is a '.3' multiply everything by 3 etc	
5. Determine molecular formula from empirical formula	$F_{mol} = \frac{M_r}{m(F_e)} \cdot F_e$ F_{mol} = molecular formula M_r = relative molecular mass F_e = empirical formula $m(F_e)$ = empirical formula mass	This and the previous calculation are often combined together in exam questions
6. Use mole ratios to determine the number of moles of B that can be made from A	$n(B) = n(A) \cdot \frac{\text{number of B in equation}}{\text{number of A in equation}}$	- The second term in this equation is the mole ratio - You must use a fully balanced equation - This is the central step in many stoichiometry calculations
7. Calculate theoretical yield	- Use Calculation 6 to determine the expected quantity of product in moles - Use a rearranged Calculation 3 to determine the expected mass.	

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Table of Contents Ib Hl Chemistry Data Let 2014

1. Understanding the eBook Ib Hl Chemistry Data Let 2014
 - The Rise of Digital Reading Ib Hl Chemistry Data Let 2014
 - Advantages of eBooks Over Traditional Books
2. Identifying Ib Hl Chemistry Data Let 2014
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ib Hl Chemistry Data Let 2014
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ib Hl Chemistry Data Let 2014

- Personalized Recommendations
 - Ib Hl Chemistry Data Let 2014 User Reviews and Ratings
 - Ib Hl Chemistry Data Let 2014 and Bestseller Lists
5. Accessing Ib Hl Chemistry Data Let 2014 Free and Paid eBooks
 - Ib Hl Chemistry Data Let 2014 Public Domain eBooks
 - Ib Hl Chemistry Data Let 2014 eBook Subscription Services
 - Ib Hl Chemistry Data Let 2014 Budget-Friendly Options
 6. Navigating Ib Hl Chemistry Data Let 2014 eBook Formats
 - ePub, PDF, MOBI, and More
 - Ib Hl Chemistry Data Let 2014 Compatibility with Devices
 - Ib Hl Chemistry Data Let 2014 Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ib Hl Chemistry Data Let 2014
 - Highlighting and Note-Taking Ib Hl Chemistry Data Let 2014
 - Interactive Elements Ib Hl Chemistry Data Let 2014
 8. Staying Engaged with Ib Hl Chemistry Data Let 2014
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ib Hl Chemistry Data Let 2014
 9. Balancing eBooks and Physical Books Ib Hl Chemistry Data Let 2014
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ib Hl Chemistry Data Let 2014
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Ib Hl Chemistry Data Let 2014
 - Setting Reading Goals Ib Hl Chemistry Data Let 2014
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Ib Hl Chemistry Data Let 2014

- Fact-Checking eBook Content of Ib Hl Chemistry Data Let 2014
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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