

## 12

## STOICHIOMETRY

## Practice Problems

In your notebook, solve the following problems.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate,  $\text{KClO}_3(\text{s})$ . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from  $3.8 \times 10^{23}$  molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

## SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



# Guided Reading And Study Workbook Chapter 12

## Stoichiometry Answers

**James Joseph Gallagher**



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## **Table of Contents Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers**

1. Understanding the eBook Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Personalized Recommendations
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers User Reviews and Ratings
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers and Bestseller Lists
5. Accessing Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers Free and Paid eBooks
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers Public Domain eBooks
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers eBook Subscription Services
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers Budget-Friendly Options

6. Navigating Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers Compatibility with Devices
  - Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Interactive Elements Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
8. Staying Engaged with Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Setting Reading Goals Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 12 Stoichiometry Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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