

## 12

## STOICHIOMETRY

## Practice Problems

In your notebook, solve the following problems.

## SECTION 12.1 THE ARITHMETIC OF EQUATIONS

Use the 3-step problem-solving approach you learned in Chapter 1.

1. An apple pie needs 10 large apples, 2 cups (top and bottom), and 1 tablespoon of cinnamon. Write a balanced equation that fits this situation. How many apples are needed to make 20 pies?
2. Two moles of potassium chloride and three moles of oxygen are produced from the decomposition of two moles of potassium chlorate,  $\text{KClO}_3(\text{s})$ . Write the balanced equation. How many moles of oxygen are produced from 12 moles of potassium chlorate?
3. Using the equation from problem 2, how many moles of oxygen are produced from 14 moles of potassium chlorate?
4. Two molecules of hydrogen react with one molecule of oxygen to produce two molecules of water. How many molecules of water are produced from  $3.8 \times 10^{23}$  molecules of oxygen? How many moles of water are produced from 22.5 moles of oxygen?

## SECTION 12.2 CHEMICAL CALCULATIONS

1. Calculate the number of moles of hydrogen chloride produced from 10 moles of hydrogen.



2. Calculate the number of moles of chlorine needed to form 14 moles of iron(II) chloride.



3. Calculate the number of grams of nitrogen dioxide that are produced from 4 moles of nitric oxide.



4. Calculate the mass of oxygen produced from the decomposition of 75.0 g of potassium chlorate.



5. Calculate the mass of silver needed to react with chlorine to produce 94 g of silver chloride. Write a balanced equation first.

6. How many liters of carbon monoxide at STP are needed to react with 4.00 g of oxygen gas to produce carbon dioxide?



7. Calculate the number of liters of oxygen gas needed to produce 15.0 liters of dinitrogen monoxide. Assume all gases are at the same conditions of temperature and pressure.



# Guided Study Workbook Answers Stoichiometry

**Brendan G. Carr**



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### **Table of Contents Guided Study Workbook Answers Stoichiometry**

1. Understanding the eBook Guided Study Workbook Answers Stoichiometry
  - The Rise of Digital Reading Guided Study Workbook Answers Stoichiometry
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Study Workbook Answers Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Study Workbook Answers Stoichiometry
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Study Workbook Answers Stoichiometry

- Personalized Recommendations
  - Guided Study Workbook Answers Stoichiometry User Reviews and Ratings
  - Guided Study Workbook Answers Stoichiometry and Bestseller Lists
5. Accessing Guided Study Workbook Answers Stoichiometry Free and Paid eBooks
    - Guided Study Workbook Answers Stoichiometry Public Domain eBooks
    - Guided Study Workbook Answers Stoichiometry eBook Subscription Services
    - Guided Study Workbook Answers Stoichiometry Budget-Friendly Options
  6. Navigating Guided Study Workbook Answers Stoichiometry eBook Formats
    - ePub, PDF, MOBI, and More
    - Guided Study Workbook Answers Stoichiometry Compatibility with Devices
    - Guided Study Workbook Answers Stoichiometry Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Guided Study Workbook Answers Stoichiometry
    - Highlighting and Note-Taking Guided Study Workbook Answers Stoichiometry
    - Interactive Elements Guided Study Workbook Answers Stoichiometry
  8. Staying Engaged with Guided Study Workbook Answers Stoichiometry
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Guided Study Workbook Answers Stoichiometry
  9. Balancing eBooks and Physical Books Guided Study Workbook Answers Stoichiometry
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Guided Study Workbook Answers Stoichiometry
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Guided Study Workbook Answers Stoichiometry
    - Setting Reading Goals Guided Study Workbook Answers Stoichiometry
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Guided Study Workbook Answers Stoichiometry



- Fact-Checking eBook Content of Guided Study Workbook Answers Stoichiometry
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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