#### Experiment 14

#### Data and Calculations: Heat Effects and Calcrimetry

#### A. Specific Heat

Trial 1		Trial 2
45,393 8		45,406 =
22, 525 8	-	22(139p 8
	-	S. 68 8
LIT.798 8		48.2918
37,175 #		31,687 =
228 68 =		_22.8(8_g
21,7 %		2150 0
99.8 ~		_(0013_°c
2513 °c		2415 °C
°c		°c
~c		°c
		3/g°C
	264.	
	45,393 = 22,525 = 3998 = 39,117 = 39,117 = 22,8 18 = 21,17 = c  99,18 = c  90,18 = c  90,18 = c	45,393 * - 22,1725 * - 3,167 * - 41,173 * - 41,173 * - 41,173 * 41,1748 * - 22,8 68 * - 21,12 * c - 25,13 * c - *

#### B. Heat of Solution

Mass of calorimeter plus water

Mass of beaker

29.309 =

(continued on following page)

# **Lab Manual Chemistry Heat Effects And Calorimetry**

**RJ Shavelson** 

**Lab Manual Chemistry Heat Effects And Calorimetry:** 

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