

Ionic Salts Simulation Lab**Learning Goals:**

- Understand how an ionic salt dissolves describing a saturated solution microscopically and macroscopically.
- Calculate solubility in grams/100ml
- Determine why some salts are more soluble than others
- Design and carry out a virtual lab to test a hypothesis

Directions: Go to

http://phet.colorado.edu/simulations/sims.php?sim=Salts_and_Solubility and run the Salts and Solubility simulation. Take a few minutes to get use to how the simulation works. Then answer the questions below.

Questions:

- 1) Describe what you observe when you shake table salt into water. Test numbers of particles with different volumes.
 - a. When the salt dissolves, which element becomes the anion (the negatively charged atom) and which one becomes the cation (the positively charged atom)?
 - b. What is the relationship between amount of particles and amount of water? (direct or indirect).
 - c. At 5.00×10^{-23} what is the point of saturation (point where you can't dissolve any more salt)?
- 2) Fill in the chart below to see if the ratio of particles to water is constant at the saturation point.

Particles	Water Volume	Ratio (particles/volume)

- a. Calculate the constant below (show your work)

- 3) Test a variety of slightly soluble salts, noting the relationship between numbers of each ion pair. Fill in the chart below to help you see a relationship.
 - a. How many particles of each element need to be present before the salt stops dissolving (point of saturation)?

Lab 30 Solubility Of A Salt Answers

Joacim Rocklöv



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Table of Contents Lab 30 Solubility Of A Salt Answers

1. Understanding the eBook Lab 30 Solubility Of A Salt Answers
 - The Rise of Digital Reading Lab 30 Solubility Of A Salt Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lab 30 Solubility Of A Salt Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lab 30 Solubility Of A Salt Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lab 30 Solubility Of A Salt Answers
 - Personalized Recommendations
 - Lab 30 Solubility Of A Salt Answers User Reviews and Ratings
 - Lab 30 Solubility Of A Salt Answers and Bestseller Lists
5. Accessing Lab 30 Solubility Of A Salt Answers Free and Paid eBooks
 - Lab 30 Solubility Of A Salt Answers Public Domain eBooks
 - Lab 30 Solubility Of A Salt Answers eBook Subscription Services
 - Lab 30 Solubility Of A Salt Answers Budget-Friendly Options
6. Navigating Lab 30 Solubility Of A Salt Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Lab 30 Solubility Of A Salt Answers Compatibility with Devices
 - Lab 30 Solubility Of A Salt Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lab 30 Solubility Of A Salt Answers
 - Highlighting and Note-Taking Lab 30 Solubility Of A Salt Answers
 - Interactive Elements Lab 30 Solubility Of A Salt Answers
 8. Staying Engaged with Lab 30 Solubility Of A Salt Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lab 30 Solubility Of A Salt Answers
 9. Balancing eBooks and Physical Books Lab 30 Solubility Of A Salt Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lab 30 Solubility Of A Salt Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Lab 30 Solubility Of A Salt Answers
 - Setting Reading Goals Lab 30 Solubility Of A Salt Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Lab 30 Solubility Of A Salt Answers
 - Fact-Checking eBook Content of Lab 30 Solubility Of A Salt Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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