

Procedure

You will use the simulation on the following website to virtually measure and record all voltages for this experiment.

<http://web.mst.edu/~gbert/Electro/Electrochem.html>

Once you have gone to the website above you should see the interface below.



- You will be able to adjust the electrodes and solutions using the drop down menus on the right and left. The simulator will only allow you to use a solution of the metal salt corresponding to the metal electrode placed in the beaker. (For example: Copper(II) nitrate must be used with the copper electrode.)
- You can also adjust the concentration of each of the metal salt solutions in the beakers by typing the value in the box on either side of the simulator.
- To measure the voltage of any cell that you have constructed, simply click **Measure Cell Voltage** on the bottom left of the simulator. If there is a problem with the cell, the simulator will identify any errors in your construction.
- You will not be using the **New Problem** link in the or "Level" drop down menu in the lower right for this assignment.

Part A

1. You will be measuring the cell potentials for all the possible combinations of each of the following electrodes: Ag/Ag^+ , Cu/Cu^{2+} , Ni/Ni^{2+} , Pb/Pb^{2+} , and Zn/Zn^{2+} . Construct the cells so that the metal salt solution in each beaker has a concentration of 0.10 M and a positive voltage is read on the voltmeter when the cell potential is measured.
2. Place a silver electrode in the beaker on the left using the top drop down menu and then select silver nitrate in the second drop down menu and type in a concentration of 0.10 into the box at the bottom.

Lab Manual For Chem 1212l

**George Chapman Caldwell, Abram
Adams Breneman**



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CHEM 1212 LAB, KSU, Huck Benjamin Huck, N. J. Kirk, 2014-01-01 Chemistry Lab Manual **Chemistry 1212 L Laboratory Manual**, 2015-12 *Laboratory Experiments in General Chemistry* Alan M. Gabrielli, Jack Duff, 2006
Laboratory Manual for Chemical and Bacterial Analysis of Water and Sewage Frank Richard Theroux, Edward F. Eldridge, Walter LeRoy Mallmann, 1936 **A Laboratory Manual for the Chemical Analysis of Water and Sewage** Edward F. Eldridge, Frank Richard Theroux, 1935 *Course and Curriculum Improvement Projects: Mathematics, Science, Social Sciences* National Science Foundation (U.S.), 1970 **Journal of the American Medical Association**, 1928
Includes proceedings of the association papers read at the annual sessions and lists of current medical literature British Medical Journal, 1908 **Boys' Life**, 1932-12 Boys Life is the official youth magazine for the Boy Scouts of America
Published since 1911 it contains a proven mix of news nature sports history fiction science comics and Scouting
Analytical Methods for a Textile Laboratory American Association of Textile Chemists and Colorists, 1968
Separations Chemistry Fedor Macásek, James D. Navratil, 2016-06-06 Separation of chemical species is a gate to final success of synthesis and preparation of compounds in pure and defined state Variability of natural and artificial mixtures to be treated is enormous Task of chemistry is to separate components of homogeneous mixtures the gaseous and liquid solutions The book concentrates on understanding the basic philosophies of both equilibrium and nonequilibrium chemical thermodynamics and engineering performance that lay in principle of separation technique such as distillation crystallization centrifugation sorption membrane separations chromatography and liquid liquid extraction Specific phenomena connected with photochemical separation isotope composition and radioactivity are discussed as well The book is written for advanced students of chemistry having the knowledge of physical chemistry Calculation examples are based on the international system of units Unique list of over 1 300 full references covers scientific literature of the eighteenth to the twenty first centuries **The Journal of Industrial and Engineering Chemistry**, 1918 Journal of Industrial and Engineering Chemistry, 1918 **Laboratory Safety for Chemistry Students** Robert H. Hill, Jr., David C. Finster, 2016-03-28 Provides knowledge and models of good practice needed by students to work safely in the laboratory as they progress through four years of undergraduate laboratory work Aligns with the revised safety instruction requirements from the ACS Committee on Professional Training 2015 Guidelines and Evaluation Procedures for Bachelor s Degree Programs Provides a systematic approach to incorporating safety and health into the chemistry curriculum Topics are divided into layers of progressively more advanced and appropriate safety issues so that some topics are covered 2 3 times at increasing levels of depth Develops a strong safety ethic by continuous reinforcement of safety to recognize assess and manage laboratory hazards and to plan for response to laboratory emergencies Covers a thorough exposure to chemical health and safety so that students will have the proper education and training when they enter the workforce or graduate school **Manual of Introductory**

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