

Chapter 4: Supramolecular Problems - Answers

1. Assign the oxidation number to each atom in the following elements, molecules and ionic substances:

- a) CO_2 b) CO c) HNO_3 d) H_2SO_4 e) H_2SO_3 f) H_2
 $\text{C} = +4$ $\text{C} = +2$ $\text{N} = +5$ $\text{S} = +6$ $\text{S} = +4$ $\text{H}_2 = 0$
 $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{H}_2 = 0$
- g) Fe h) FeSO_4 i) CO_2 j) H_2O k) HNO_3 l) O_2
 $\text{Fe} = 0$ $\text{Fe} = +2$ $\text{C} = +4$ $\text{H} = +1$ $\text{N} = +5$ $\text{O}_2 = 0$
 $\text{S} = +6$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O}_2 = 0$

2. Assign the oxidation number to each atom in the following ions:

- a) NO^+ b) CO_3^{2-} c) HNO_2 d) SO_4^{2-} e) NO_2^+ f) HCO_3^-
 $\text{N} = +3$ $\text{C} = +4$ $\text{N} = +3$ $\text{S} = +6$ $\text{N} = +3$ $\text{C} = +4$
 $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$ $\text{O} = -2$

3. Write the following reactions identifying if it is an oxidation-reduction reaction. If it is a redox reaction identify the oxidizing agent, the reducing agent, the substance being oxidized and the substance being reduced. Identify the oxidation number change to help to support these identifications.



Redox rxn. Cu oxidized (oxidizing agent). H_2 is $0 \rightarrow +1$
 HCl reduced (oxidizing agent). H is $+1 \rightarrow 0$



Not a redox rxn. Cl is $+1 \rightarrow 0$ but both reactants and product



Redox rxn. H_2 oxidized (oxidizing agent). H is $0 \rightarrow +1$
 H_2O reduced (oxidizing agent). H is $+1 \rightarrow 0$



Not a redox reaction. No change in ox. #



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General College Chemistry Charles William Keenan, Donald C. Kleinfelter, Jesse Hermon Wood, 1980

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