Chemistry: Balancing Chemical Equations

<u>Directions</u>: First, balance each of the chemical equations below. Then, classify each reaction as **synthesis**, **decomposition**, **single-replacement**, or **double-replacement**. To earn full credit, write the words out when classifying.

Balance the equation...

- _2_Sb + _3_Cl₂ → _2_SbCl₃
- 2. Mg + ___O; → __2_MgO
- 3. ____ CaCl₂ → ___ Ca + ___ Cl₂
- 2 NaClO₂ → 2 NaCl + 3 O₂
- Fe + 2 HCl → FeCl₂ + H₂
- 6. CuO + H₂ → Cu + H₂O
- 7. 2 Al + 3 H₂SO₄ → Al₂(SO₄)₁ + 3 H₂
- MgBr₂ + ___ Cl₂ → ___ MgCl₂ + ___ Br₂
- 9. ____SnO₂ + __2 C → ____Sn + __2 CO
- 10. $Pb(NO_3)_2 + ___ H_2S \rightarrow ___ PbS + __2 HNO_3$
- 11. __2_HgO → __2_Hg + ____O;
- 12. __2_KClO₃ → __2_KCl + __3_O₂
- 13. $N_2 + 3 H_2 \rightarrow 2 NH_3$
- 14. __2 NaBr + ___ Cl₂ → __2 NaCl + ___ Br₂
- 15. ____ Zn + __2_ AgNO₃ → ____ Zn(NO₃)₂ + __2_ Ag
- 16. ____ Sn + __2 Cl₂ → ____ SnCl₄
- 17. ____ Ba(OH)₂ → ____ BaO + ____ H₂O

...and classify it.

synthesis

synthesis

decomposition

decomposition

single replacement

single replacement

single replacement

single replacement

single replacement

double replacement

decomposition

decompostion

synthesis

single replacement

single replacement

synthesis

decomposition

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Martha Mackin

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