

Calorimetry: Heat of Neutralization

Prepared by: Dongling Fei
Manatee Community College

Purpose: Determine the calorimeter constant for a calorimeter. Determine the enthalpy of neutralization for the reaction of a strong acid and a strong base.

Background: When a system absorbs heat (q) its temperature changes by ΔT . Experimental measurements demonstrate that the heat absorbed by a system and its corresponding temperature change are directly proportional. $q = C \times \Delta T$. C is called the **heat capacity**, which is often defined as the quantity of heat required to change its temperature by 1°C .

$$C = q / \Delta T = \text{J} / ^\circ\text{C}$$

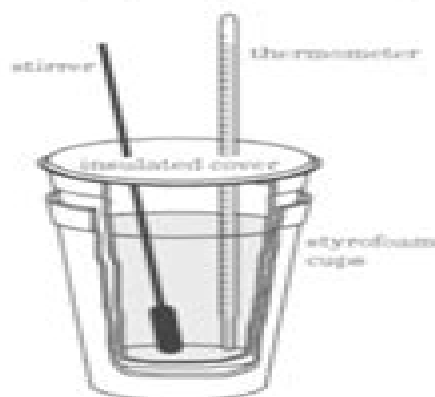
Heat capacity is an extensive property; it depends on the amount of matter being heated. The measure of intrinsic capacity of a substance to absorb heat is called **specific heat capacity** (C_s), the amount of heat required to raise the temperature of 1 gram of the substance by 1°C .

$$C_s = q / (\Delta T \times m)$$

The C_s of H_2O is $4.184 \text{ J/g}^\circ\text{C}$. For the case where all the heat is used to heat water, the amount of heat is calculated by the relation

$$q = C_s \times m_{\text{water}} \times \Delta T$$

The enthalpy change for a chemical reaction, ΔH_{rxn} is also called the **enthalpy of reaction** or **heat of reaction**, which is also an extensive property. For many aqueous reactions, ΔH_{rxn} can be measured simply using a coffee-cup calorimeter



Heat Of Neutralization Lab Manual

Albert A Gayle



Heat Of Neutralization Lab Manual:

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