

## Chem Exam Study Guide

### Experiment 1: Polarity and Electrolytes

**Know how to convert from one SI unit to another (Example: meters to nanometers)**

meters to nanometers

$$1 \text{ m} = 1000000000 \text{ nm}$$

Nano = Nine

Milliliter to microliter

$$1 \text{ mL} = 1000 \text{ microL}$$

Liter to Milliliter

$$1 \text{ L} = 1000 \text{ mL}$$

**Know the basic intermolecular forces and their relative strengths (London dispersion, dipole-dipole, hydrogen bonding)**

Strength from greatest to least greatest

#### Hydrogen Bonding

- 4 kJ to 50 kJ
- Chemical bond
- interaction involving a hydrogen atom located between a pair of other atoms having a high affinity for electrons
- Nitrogen, Oxygen, Nitrogen, Fluorine, Water

#### Dipole-Dipole Force

- 5 kJ to 20 kJ
- Attractive forces between the positive end of one polar molecule and the negative end of another polar molecule
- Hydrogen Chloride, Carboxylic Acids, Amino Acids

#### London Dispersion Force

- a temporary attractive force that results when the electrons in two adjacent atoms occupy positions that make the atoms form temporary dipoles
- Helium, Halogens, Chlorine, Bromine, Iodine, Neon, Argon

# Lab Chem Study Guide 2015 Final Exam

**M Carnoy**



## **Lab Chem Study Guide 2015 Final Exam:**

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