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### Solubility Curve Worksheet and Lab



You'll notice that for most substances, solubility increases as temperature increases. As discussed earlier in solutions involving liquids and solids typically more solute can be dissolved at higher temperatures. Can you find any exceptions on the graph? \_\_\_\_\_

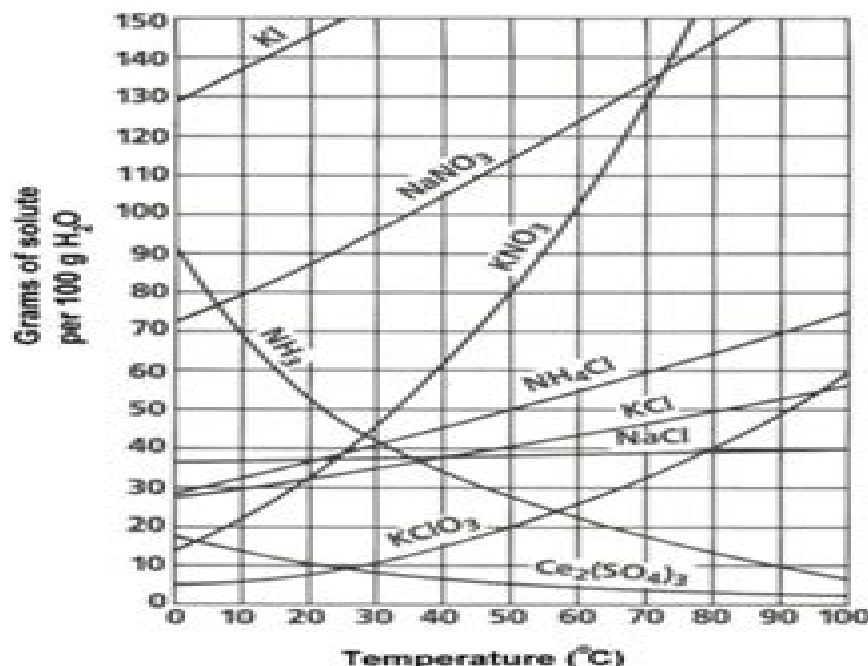
Here's an example of how to read the graph. Find the curve for  $\text{KClO}_3$ .

At  $30^\circ\text{C}$  approximately 10g of  $\text{KClO}_3$  will dissolve in 100g of water. If the temperature is increased to  $80^\circ\text{C}$ , approximately \_\_\_\_\_ of the substance will dissolve in 100g (or 100mL) of water.

**Directions:** Use the graph to answer the following questions. REMEMBER UNITS!

1) What mass of solute will dissolve in 100mL of water at the following temperatures?

- $\text{KNO}_3$  at  $70^\circ\text{C}$  = \_\_\_\_\_
- $\text{NaCl}$  at  $100^\circ\text{C}$  = \_\_\_\_\_
- $\text{NH}_4\text{Cl}$  at  $90^\circ\text{C}$  = \_\_\_\_\_
- Which of the above three substances is most soluble in water at  $15^\circ\text{C}$ . = \_\_\_\_\_



# Lab Solubility Data Sheet Answer Key

**KJ Lindholm-Leary**



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**Solubilities of Inorganic and Organic Compounds** Atherton Seidell, 1958

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