

temperature:

- a measure of the average kinetic energy of the particles in a sample of matter
- does not depend on the amount of matter in the sample
- symbol is T; unit is °C

heat:

- total amount of kinetic energy that flows because of a difference in temperature
- depends on mass of sample
- symbol is Q; unit is J (1 J = 4.18 cal)

Kinetic energy is energy in motion

Potential energy is store

- Potential energy is hiding and cannot be measured
- Only changes in P.E. can be measured.

specific heat capacity:

- amount of energy required to raise the temperature of 1 gram of substance 1 °C
- symbol is c; unit is J/g°C

$$\begin{array}{c}
 \text{mass} \swarrow \\
 Q = m \times C \times \Delta t \\
 \swarrow \quad \quad \quad \nwarrow \\
 \text{heat} \quad \quad \quad \text{specific heat} \quad \quad \quad \text{change in temperature}
 \end{array}$$

When heat (Q) is absorbed by a system, part of it (C) goes into storage as potential energy and part of it is used to make the molecules move around more, raising the temperature (Δt).

Why does sand get hotter in the day and colder at night than the water?

Gpb Physics Note Taking Guide Episode 1301

TD Snyder



Gpb Physics Note Taking Guide Episode 1301:

This book delves into Gpb Physics Note Taking Guide Episode 1301. Gpb Physics Note Taking Guide Episode 1301 is a crucial topic that must be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gpb Physics Note Taking Guide Episode 1301, encompassing both the fundamentals and more intricate discussions.

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 - Chapter 4: Gpb Physics Note Taking Guide Episode 1301 in Specific Contexts
 - Chapter 5: Conclusion
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