

Ionic Salts Simulation Lab**Learning Goals:**

- Understand how an ionic salt dissolves describing a saturated solution microscopically and macroscopically.
- Calculate solubility in grams/100ml
- Determine why some salts are more soluble than others
- Design and carry out a virtual lab to test a hypothesis

Directions: Go to

http://phet.colorado.edu/simulations/sims.php?sim=Salts_and_Solubility and run the Salts and Solubility simulation. Take a few minutes to get use to how the simulation works. Then answer the questions below.

Questions:

- 1) Describe what you observe when you shake table salt into water. Test numbers of particles with different volumes.
 - a. When the salt dissolves, which element becomes the anion (the negatively charged atom) and which one becomes the cation (the positively charged atom)?
 - b. What is the relationship between amount of particles and amount of water? (direct or indirect).
 - c. At 5.00×10^{-23} what is the point of saturation (point where you can't dissolve any more salt)?
- 2) Fill in the chart below to see if the ratio of particles to water is constant at the saturation point.

Particles	Water Volume	Ratio (particles/volume)

- a. Calculate the constant below (show your work)

- 3) Test a variety of slightly soluble salts, noting the relationship between numbers of each ion pair. Fill in the chart below to help you see a relationship.
 - a. How many particles of each element need to be present before the salt stops dissolving (point of saturation)?

Lab 30 Solubility Of A Salt Answers

Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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web answers 1 H_2O 2 H_2O 3 H_2O 4 H_2O 5 H_2O 6 H_2O 7 H_2O 8 H_2O 9 H_2O 10 H_2O 11 H_2O 12 H_2O 13 H_2O 14 H_2O 15 H_2O 16 H_2O 17 H_2O 18 H_2O 19 H_2O 20 H_2O 21 H_2O 22 H_2O 23 H_2O 24 H_2O 25 H_2O 26 H_2O 27 H_2O 28 H_2O 29 H_2O 30 H_2O 31 H_2O 32 H_2O 33 H_2O 34 H_2O 35 H_2O 36 H_2O 37 H_2O 38 H_2O 39 H_2O 40 H_2O 41 H_2O 42 H_2O 43 H_2O 44 H_2O 45 H_2O 46 H_2O 47 H_2O 48 H_2O 49 H_2O 50 H_2O 51 H_2O 52 H_2O 53 H_2O 54 H_2O 55 H_2O 56 H_2O 57 H_2O 58 H_2O 59 H_2O 60 H_2O 61 H_2O 62 H_2O 63 H_2O 64 H_2O 65 H_2O 66 H_2O 67 H_2O 68 H_2O 69 H_2O 70 H_2O 71 H_2O 72 H_2O 73 H_2O 74 H_2O 75 H_2O 76 H_2O 77 H_2O 78 H_2O 79 H_2O 80 H_2O 81 H_2O 82 H_2O 83 H_2O 84 H_2O 85 H_2O 86 H_2O 87 H_2O 88 H_2O 89 H_2O 90 H_2O 91 H_2O 92 H_2O 93 H_2O 94 H_2O 95 H_2O 96 H_2O 97 H_2O 98 H_2O 99 H_2O 100 H_2O the enthalpy of reaction is ΔH prod ΔH react the reactant is at a higher potential than the product when the reaction occurs heat is released stabilizing the system and the system moves to a lower energy potential 3

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web thermochemical equations quiz this online quiz is intended to give you extra practice in performing thermochemical calculations with a variety of reactions including limiting reagents and percent yield options this quiz aligns with the following ngss standard s hs ps1 4 hs ps3 1

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web in this set of practice questions we will summarize the main concepts of thermochemistry such as the relationship between internal energy work and heat exothermic and endothermic process heat capacity constant pressure calorimetry constant volume calorimetry the enthalpy the standard enthalpies of formation and their use in