

Name: _____ Date: _____ Lab Section: _____

Experiment 2: Molecular Models REPORT SHEET

Pre-lab Study Questions:

- Objectives of the experiment (in your own words):
- What kinds of electrons form the covalent bonds in molecules?
- Is a polar molecule symmetrical or asymmetrical?

FORMULA OF MOLECULE	LEWIS STRUCTURE	3-D SHAPE w/ lone pairs	MOLECULAR GEOMETRY	MOLECULAR POLARITY
example: <chem>CH4</chem>	<pre> H H - C - H H </pre>		tetrahedral	nonpolar
<chem>H2O</chem>	<pre> .. O / \ H H </pre>	<pre> H .. P \ / O / \ H H </pre>	tetrahedral	polar
<chem>H2O2</chem>	<pre> H : O : H .. O / \ H H </pre>	<pre> H .. O .. H \ / \ / O O / \ / \ H H H H </pre>	tetrahedral	nonpolar
<chem>C2H6</chem>				
<chem>C2H4</chem>	<pre> H C H \ / \ / C=C / \ / \ H C H </pre>		trigonal planar	non polar
<chem>C2H2</chem>				
<chem>CO2</chem>	<pre> O = C = O </pre>		Linear	Non polar
<chem>H2O2</chem>				
<chem>CH3Cl</chem>				

Lab 22 Models Molecular Compounds Answer

G Psacharopoulos



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