

Experiment 14

Data and Calculations: Heat Effects and Calorimetry

A. Specific Heat

	Trial 1		Trial 2
Mass of stoppered test tube plus metal	<u>45.393</u> g	→	<u>45.406</u> g
Mass of test tube and stopper	<u>22.525</u> g	→	<u>22.539</u> g
Mass of calorimeter	<u>3.978</u> <u>47.173</u> g	→	<u>8.608</u> g
Mass of calorimeter and water	<u>45.798</u> g		<u>48.295</u> g
Mass of water	<u>39.175</u> g		<u>39.687</u> g
Mass of metal	<u>22.868</u> g	→	<u>22.868</u> g
Initial temperature of water in calorimeter	<u>21.2</u> °C		<u>21.9</u> °C
Initial temperature of metal (assume 100°C unless directed to do otherwise)	<u>99.8</u> °C	→	<u>100.3</u> °C
Equilibrium temperature of metal and water in calorimeter	<u>25.3</u> °C		<u>24.5</u> °C
$\Delta t_{\text{water}} (t_{\text{final}} - t_{\text{initial}})$	_____ °C		_____ °C
Δt_{metal}	_____ °C		_____ °C
$q_{\text{H}_2\text{O}}$	_____ J		_____ J
Specific heat of the metal (Eq. 3)	_____ J/g°C		_____ J/g°C
Approximate molar mass of metal	_____		_____
Unknown metal #			<u>26Y</u>

B. Heat of Solution

Mass of calorimeter plus water	<u>56.898</u> g
Mass of beaker	<u>29.309</u> g

(continued on following page)

Lab Manual Chemistry Heat Effects And Calorimetry

AW Rasmussen



Lab Manual Chemistry Heat Effects And Calorimetry:

Reviewing **Lab Manual Chemistry Heat Effects And Calorimetry**: Unlocking the Spellbinding Force of Linguistics

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