



LS Retail – Training Manual

Version 4.0

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web answer see below explanation the rate of a chemical reaction $k = A e^{-E_a/RT}$ is given in first approximation by $k = A e^{-E_a/RT}$ where k depends by many factors as solvent geometry type of reaction E_a is the activation energy R is the universal constant of gases T is the absolute temperature

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web the energy of the reactants is shown on the energy diagram above on the right side of the energy diagram draw a horizontal line segment to indicate the energy of the products draw a vertical line segment to indicate ΔH for the reaction and label it with the correct value see attached image

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