

temperature:

- a measure of the average kinetic energy of the particles in a sample of matter
- does not depend on the amount of matter in the sample
- symbol is T; unit is °C

heat:

- total amount of kinetic energy that flows because of a difference in temperature
- depends on mass of sample
- symbol is Q; unit is J (1 J = 4.18 cal)

Kinetic energy is energy in motion

Potential energy is store

- Potential energy is hiding and cannot be measured
- Only changes in P.E. can be measured.

specific heat capacity:

- amount of energy required to raise the temperature of 1 gram of substance 1 °C
- symbol is c; unit is J/g°C

$$\begin{array}{ccccc}
 & \text{mass} & & & \\
 & \downarrow & & & \\
 \text{heat} \nearrow & Q = m \times C \times \Delta t & \nwarrow & & \text{change in temperature} \\
 & \text{specific heat} & & &
 \end{array}$$

When heat (Q) is absorbed by a system, part of it (C) goes into storage as potential energy and part of it is used to make the molecules move around more, raising the temperature (Δt).

Why does sand get hotter in the day and colder at night than the water?

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DP Hallahan



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Table of Contents Gpb Physics Note Taking Guide Episode 1301

1. Understanding the eBook Gpb Physics Note Taking Guide Episode 1301
 - The Rise of Digital Reading Gpb Physics Note Taking Guide Episode 1301
 - Advantages of eBooks Over Traditional Books
2. Identifying Gpb Physics Note Taking Guide Episode 1301
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gpb Physics Note Taking Guide Episode 1301
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gpb Physics Note Taking Guide Episode 1301
 - Personalized Recommendations
 - Gpb Physics Note Taking Guide Episode 1301 User Reviews and Ratings
 - Gpb Physics Note Taking Guide Episode 1301 and Bestseller Lists
5. Accessing Gpb Physics Note Taking Guide Episode 1301 Free and Paid eBooks
 - Gpb Physics Note Taking Guide Episode 1301 Public Domain eBooks
 - Gpb Physics Note Taking Guide Episode 1301 eBook Subscription Services
 - Gpb Physics Note Taking Guide Episode 1301 Budget-Friendly Options

6. Navigating Gpb Physics Note Taking Guide Episode 1301 eBook Formats
 - ePub, PDF, MOBI, and More
 - Gpb Physics Note Taking Guide Episode 1301 Compatibility with Devices
 - Gpb Physics Note Taking Guide Episode 1301 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gpb Physics Note Taking Guide Episode 1301
 - Highlighting and Note-Taking Gpb Physics Note Taking Guide Episode 1301
 - Interactive Elements Gpb Physics Note Taking Guide Episode 1301
8. Staying Engaged with Gpb Physics Note Taking Guide Episode 1301
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gpb Physics Note Taking Guide Episode 1301
9. Balancing eBooks and Physical Books Gpb Physics Note Taking Guide Episode 1301
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gpb Physics Note Taking Guide Episode 1301
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gpb Physics Note Taking Guide Episode 1301
 - Setting Reading Goals Gpb Physics Note Taking Guide Episode 1301
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gpb Physics Note Taking Guide Episode 1301
 - Fact-Checking eBook Content of Gpb Physics Note Taking Guide Episode 1301
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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