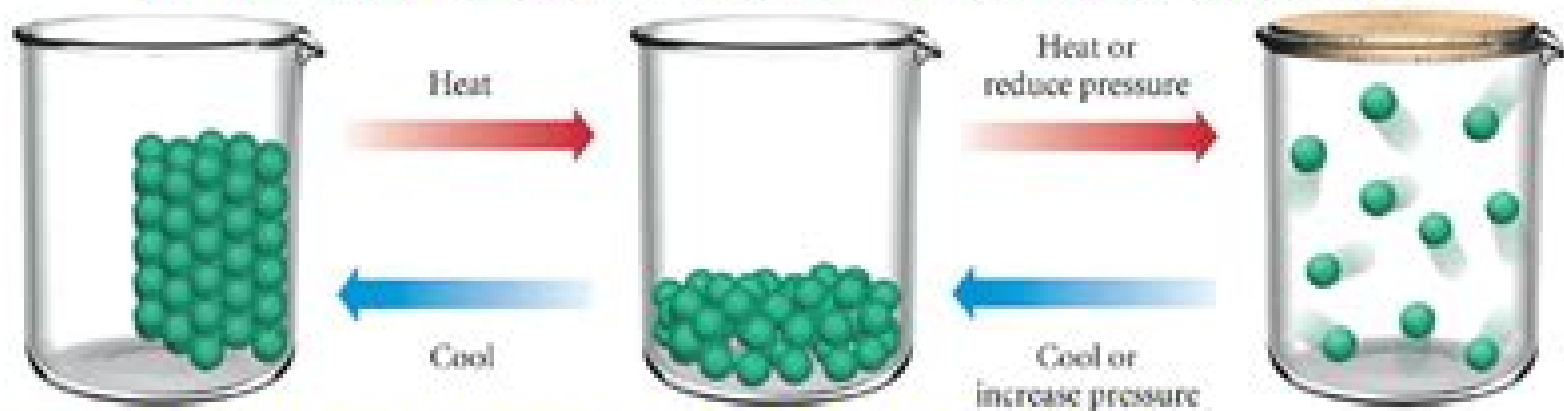


# Chapter 11: Liquids and Solids

## A. Intermolecular Forces

1. The addition of heat changes the state of a compound from solid to liquid to gas as kinetic and potential energy are added to the molecules
  - a. Kinetic Energy involves the motion of the molecules
  - b. Potential Energy involves how far apart the molecules are



2. Intermolecular Forces hold the molecules together
  - a. Ion-Dipole Interactions:  $\text{Na}^+\text{Cl}^-$  interactions with water
  - b. Dipole-Dipole Attractions: interaction of polar molecules
  - c. Hydrogen Bonding: special dipole-dipole attraction in molecules where H is bonded to an electronegative atom (O, N, F, etc...)
    - i. Small size of H lets the dipoles get very close
    - ii. Large electronegativity difference makes attraction great

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**Camilla Rothe**



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### **Table of Contents Grade11 Experiment Intermolecular Force**

1. Understanding the eBook Grade11 Experiment Intermolecular Force
  - The Rise of Digital Reading Grade11 Experiment Intermolecular Force
  - Advantages of eBooks Over Traditional Books
2. Identifying Grade11 Experiment Intermolecular Force
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade11 Experiment Intermolecular Force
  - User-Friendly Interface
4. Exploring eBook Recommendations from Grade11 Experiment Intermolecular Force
  - Personalized Recommendations
  - Grade11 Experiment Intermolecular Force User Reviews and Ratings
  - Grade11 Experiment Intermolecular Force and Bestseller Lists

5. Accessing Grade11 Experiment Intermolecular Force Free and Paid eBooks
  - Grade11 Experiment Intermolecular Force Public Domain eBooks
  - Grade11 Experiment Intermolecular Force eBook Subscription Services
  - Grade11 Experiment Intermolecular Force Budget-Friendly Options
6. Navigating Grade11 Experiment Intermolecular Force eBook Formats
  - ePub, PDF, MOBI, and More
  - Grade11 Experiment Intermolecular Force Compatibility with Devices
  - Grade11 Experiment Intermolecular Force Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade11 Experiment Intermolecular Force
  - Highlighting and Note-Taking Grade11 Experiment Intermolecular Force
  - Interactive Elements Grade11 Experiment Intermolecular Force
8. Staying Engaged with Grade11 Experiment Intermolecular Force
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade11 Experiment Intermolecular Force
9. Balancing eBooks and Physical Books Grade11 Experiment Intermolecular Force
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Grade11 Experiment Intermolecular Force
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Grade11 Experiment Intermolecular Force
  - Setting Reading Goals Grade11 Experiment Intermolecular Force
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Grade11 Experiment Intermolecular Force
  - Fact-Checking eBook Content of Grade11 Experiment Intermolecular Force
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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